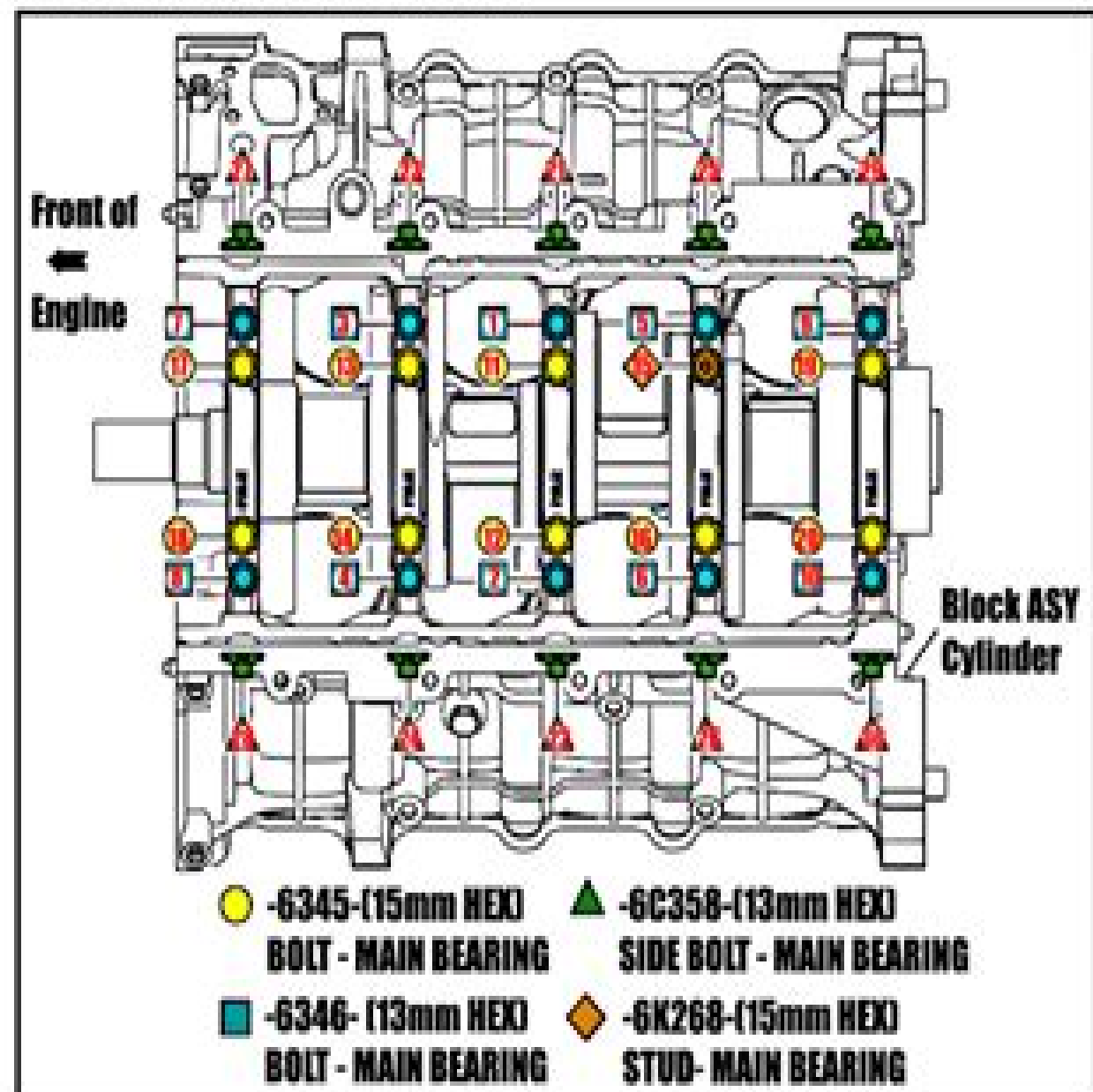


MAIN BEARING CAP TORQUE

M-6007-A50HA5.0L Engine Assembly



1. Torque Fasteners №1 - №20
in Sequence 20 Nm

2. Torque Outer Fasteners №1-№10
in Sequence 40 Nm

3. Torque Inner Fasteners №11-№20
in Sequence 65 Nm

4. Rotate Fasteners №1-№20
in Sequence Clockwise 90

5. Torque Side Bolts №21 - №30
in Sequence 10 Nm

6. Torque Side Bolts №21 - №30
in Sequence 30 Nm

7. Rotate Side Bolts №21-№30
in Sequence Clockwise 60.

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web the relative electronegativities of the bonded atoms is h c cl and so the bond moments all point toward the cl end of the molecule and sum to yield a considerable dipole moment the molecules are relatively polar for molecules of high symmetry such as bf₃ trigonal planar ch₄ tetrahedral pf₅ trigonal bipyramidal and sf₆

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a net ionic equation is the most accurate representation of the actual chemical process that occurs writing these equations requires a familiarity with solubility rules acid base reactivity

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apr 25 2017 this is understood when you consider the two chemicals appear unchanged before and after the reaction

therefore the net ionic equation is $2\text{Cl}^- + 2\text{Ag}^+ \rightarrow 2\text{AgCl}$

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ions that remain essentially unchanged during a reaction are called spectator ions this means you can ignore them when you write the ionic equation you only need to model how the solid

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the first step to writing a net ionic equation is balancing the chemical equation present let s use the reaction between sodium chloride and silver nitrate as an example next we write the

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highlights learning objectives by the end of this section you will be able to derive chemical equations from narrative descriptions of chemical reactions write and balance chemical

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the net ionic equation is the chemical equation that shows only those elements compounds and ions that are directly involved in the chemical reaction notice that in writing the net ionic

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what remains when the spectator ions are removed is called the net ionic equation which represents the actual chemical change occurring between the ionic compounds cl aq

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this net ionic equation indicates that solid silver chloride may be produced from dissolved chloride and silver i ions regardless of the source of these ions these molecular and

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write the overall chemical equation the complete ionic equation and the net ionic equation for the reaction of aqueous silver fluoride with aqueous sodium phosphate to give solid silver

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the complete ionic equation for this reaction is as follows equation 4 13 2ag aq 2f aq 2nh₄ aq cr₂o₇²⁻ aq ag₂cr₂o₇ s 2nh₄ aq 2f aq because two nh₄ aq

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