

Key

Kinetics Practice Problems

Ex. 1: Consider the following reaction, $\text{NH}_4^+(\text{aq}) + \text{NO}_2^-(\text{aq}) \rightarrow \text{N}_2(\text{g}) + \text{H}_2\text{O}(\text{l})$, at 25°C.

- a. Determine the rate law for the reaction.

$$\text{rate} = k [\text{NH}_4^+]^x [\text{NO}_2^-]^y$$

$[\text{NH}_4^+]$ (in M)	$[\text{NO}_2^-]$ (in M)	Initial Rate (in M/s)
0.24	0.10	7.2×10^{-6}
0.12	0.10	3.6×10^{-6}
0.12	0.15	5.4×10^{-6}

$$\frac{\text{rate}_1}{\text{rate}_2} = \frac{k [\text{NH}_4^+]_1^x [\text{NO}_2^-]_1^y}{k [\text{NH}_4^+]_2^x [\text{NO}_2^-]_2^y} = \frac{\cancel{k} [0.24\text{M}]^x [\cancel{0.10\text{M}}]^y}{\cancel{k} [0.12\text{M}]^x [0.10\text{M}]^y} = \frac{7.2 \times 10^{-6} \text{ M/s}}{3.6 \times 10^{-6} \text{ M/s}} \rightarrow (2.0)^x = 2.0 \rightarrow x=1$$

$$\frac{\text{rate}_3}{\text{rate}_2} = \frac{k [\text{NH}_4^+]_3^x [\text{NO}_2^-]_3^y}{k [\text{NH}_4^+]_2^x [\text{NO}_2^-]_2^y} = \frac{k [0.12\text{M}]^x [0.15\text{M}]^y}{\cancel{k} [0.12\text{M}]^x [0.10\text{M}]^y} = \frac{5.4 \times 10^{-6} \text{ M/s}}{3.6 \times 10^{-6} \text{ M/s}} \rightarrow (1.5)^y = 1.5 \rightarrow y=1$$

$$\text{rate} = k [\text{NH}_4^+] [\text{NO}_2^-]$$

- b. Determine the rate constant for the reaction.

$$k = \frac{\text{rate}_1}{[\text{NH}_4^+]_1 [\text{NO}_2^-]_1} = \frac{7.2 \times 10^{-6} \text{ M/s}}{(0.24\text{M})(0.10\text{M})} = 3.0 \times 10^{-4} \text{ M}^{-1} \text{ s}^{-1}$$

Ex. 2: At 600°C acetone (CH_3COCH_3) decomposes to form several compounds.

- a. Determine the rate law for the reaction.

$$\text{rate} = k [\text{CH}_3\text{COCH}_3]^x$$

Experiment	$[\text{CH}_3\text{COCH}_3]$ (in M)	Initial Rate (in M/s)
1	6.0×10^{-3}	5.2×10^{-5}
2	9.0×10^{-3}	7.8×10^{-5}

$$\frac{\text{rate}_2}{\text{rate}_1} = \frac{k [\text{CH}_3\text{COCH}_3]_2^x}{k [\text{CH}_3\text{COCH}_3]_1^x} = \frac{\cancel{k} [0.0090\text{M}]^x}{\cancel{k} [0.0060\text{M}]^x} = \frac{7.8 \times 10^{-5} \text{ M/s}}{5.2 \times 10^{-5} \text{ M/s}} \rightarrow (1.5)^x = 1.5 \rightarrow x=1$$

$$\text{rate} = k [\text{CH}_3\text{COCH}_3]$$

- b. Calculate the rate constant for the reaction.

$$k = \frac{\text{rate}_1}{[\text{CH}_3\text{COCH}_3]_1} = \frac{5.2 \times 10^{-5} \text{ M/s}}{0.0060\text{M}} = 8.7 \times 10^{-3} \text{ s}^{-1}$$

- c. What is the reaction order for this reaction at 600°C? 1st-order

Skill Practice 54 Answers Chemistry

Lei Shi

A red circular graphic with a gradient, appearing as a semi-circle or a partial circle, located to the right of the author's name.

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Merrill Chemistry Robert C. Smoot,Smoot,Richard G. Smith,Jack Price,1998

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