

- ④ To find the molecular formula, multiply the empirical formula by the ratio of molar mass divided by empirical mass. In this example it is $66.0/33.0$, which equals 2.
- 4 a Calcium ions are $2+$ and chloride ions are $1-$, so the formula is CaCl_2 .
 b Silver ions are $1+$ and nitrate ions are $1-$, so the formula is AgNO_3 .
 c Copper(II) is Cu^{2+} and phosphate is PO_4^{3-} . For a neutral compound you need three copper ions and two phosphate ions, making charges of $+6$ and -6 . Therefore, the formula is $\text{Cu}_3(\text{PO}_4)_2$.
 d Aluminium is $3+$ and oxide is $2-$, so aluminium oxide is Al_2O_3 .
 ⑤ There are two ways of working out formulae of ionic compounds. One way is to make sure that the total positive charge equals the total negative charge (see part c). The other way, if the charges are not the same, is to swap them round. Copper is $2+$ and phosphate $3-$, so there are three copper ions and two phosphate ions in the formula.
- 5 a As there is one manganese ion to each carbonate ion, the charges have the same numerical value. Carbonate ions are $2-$, so the manganese ions in this compound must be $2+$.
 b The charge on two vanadium ions must equal the charge on three sulfate ions. Each sulfate is -2 , so three sulfate ions have a total charge of -6 . This means that two vanadium ions have a total charge of $+6$, so each has a charge of $+3$.
- 6 a molar mass of $\text{Ca}(\text{OH})_2 = 40.1 + 2 \times (16.0 + 1.0) = 74.1 \text{ g mol}^{-1}$
 b molar mass of $\text{Al}_2(\text{SO}_4)_3 = (2 \times 27.0) + 3 \times [32.1 + (4 \times 16.0)] = 342.3 \text{ g mol}^{-1}$
 c molar mass of $\text{FeSO}_4 \cdot 7\text{H}_2\text{O} = 55.8 + 32.1 + (4 \times 16.0) + 7 \times (2.0 + 16.0) = 277.9 \text{ g mol}^{-1}$
 ⑥ In calculating molar mass, make sure that you use the relative atomic masses (the larger number), not the atomic numbers, from the periodic table. Molar mass has the units of g mol^{-1} . Relative molecular mass has no units. The numbers are the same.
- 7 The balanced equations are:
 a $\text{P}_4 + 5\text{O}_2 \rightarrow \text{P}_4\text{O}_{10}$ or $2\text{P}_2\text{O}_5$
 ⑥ Phosphorus(V) oxide exists as the dimer P_4O_{10} , but $2\text{P}_2\text{O}_5$ on the right of the equation would be acceptable.
 b $4\text{Al} + 3\text{O}_2 \rightarrow 2\text{Al}_2\text{O}_3$
 c $2\text{Mg}(\text{NO}_3)_2 \rightarrow 2\text{MgO} + 4\text{NO}_2 + \text{O}_2$
 ⑥ In part c, balance the oxygen atoms last. Remember that the number of oxygen atoms on the right-hand side of the equation must be an even number, because $\text{Mg}(\text{NO}_3)_2$ contains an even number of oxygen atoms.
 d $2\text{LiOH} + \text{H}_2\text{SO}_4 \rightarrow \text{Li}_2\text{SO}_4 + 2\text{H}_2\text{O}$
 e $2\text{Fe}^{2+} + 2\text{I}^- \rightarrow 2\text{Fe}^{3+} + \text{I}_2$
 ⑥ Make sure that this equation balances for charge. Having one Fe^{2+} ion on the left-hand side and one Fe^{3+} on the right-hand side balances the number of Fe ions, but the charge on the left would not equal the charge on the right.
 f $2\text{C}_8\text{H}_{18} + 25\text{O}_2 \rightarrow 16\text{CO}_2 + 18\text{H}_2\text{O}$ or $\text{C}_8\text{H}_{18} + 12\frac{1}{2}\text{O}_2 \rightarrow 8\text{CO}_2 + 9\text{H}_2\text{O}$
 ⑥ The only source of oxygen on the left-hand side of the equation is O_2 , so the number of oxygen atoms on the right-hand side must be an even number. Therefore, there must be an even number of H_2O molecules on the right. This requires two C_8H_{18} molecules on the left.
- 8 For equations a–c, the simplest way is to realise that the component ions of the compound combine to give the precipitate. All other ions are spectators.

Teaching Transparency Answers Key For

Elsie Higa



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Resources in Education ,1998 **Te HS&T 2007 Shrt Crs M** Holt Rinehart & Winston,2007 **American Republic to 1877 Teaching Transparency Sampler 2003** McGraw-Hill Staff,2002-09 **Resources in Vocational Education** ,1980 *Official Gazette of the United States Patent Office* United States. Patent Office,1970 **Ven Connigo!, Grades 7-12** Holt,Holt Staff,1996 Memoirs of a Certain Sansei: A Japanese American Mother and Teacher Elsie Higa,2015-01-28 For more than four decades author and teacher Elsie Higa has inspired thousands of students many of whom write complimentary notes to her In Memoirs of a Certain Sansei she shares her stories including her parenting secrets This memoir also springs from the comments of her students and her responses to them These two forces inspired Higa to combine her knowledge of parenthood with her teaching expertise in a warm and wise narrative that parents teachers students and children will learn from Without guidance from Sensei I would not be where I am today alumnus Higa set out to share parental and scholarly wisdom and she achieved far more through stories packed with illumination on life in general They are beautifully written fast paced entertaining and enlightening Any reader in search of tips for better living effective relationships and personal reward will enjoy and profit from this book Dr Doris M Ching Vice President Emeritus Student Affairs University of Hawaii **Reading Without Limits** Maddie Witter,2013-01-07 Imagine a classroom where all students are engaged in highly rigorous and fun learning every single day That classroom can be yours starting tomorrow You don t have to be a reading specialist to pick up this book Anyone who wants to dramatically improve reading achievement will find helpful suggestions You might be a third grade teacher whose students have mastered decoding and you are ready to build their comprehension Or you might be a high school science teacher whose students aren t yet reading on level with deep critical thinking This book is for you It doesn t matter whether you are a public charter private or alternative education teacher the Reading Without Limits program works in each one Along with hundreds of ready to use teaching strategies Reading Without Limits comes with a supplemental website where teachers can download even more resources for free Reading Without Limits is the first book offered in the KIPP Educator Series KIPP or the Knowledge is Power Program began in 1994 As of Fall 2012 there are 125 KIPP schools in 20 states and the District of Columbia serving nearly 40 000 students climbing the mountain to and through college Komm Mit! Holt, Rinehart and Winston Staff,1995 Chapter Resource 3 Cell Structure Biology Holt Rinehart & Winston,Holt, Rinehart and Winston Staff,2004 Math Minutes, 5th Grade, eBook Jo Fitzpatrick, Help students master essential math skills with the motivating classroom tested Math Minutes format featured in this new book It provides 100 Minutes of 10 problems each for students to complete within a one to two minute period Providing daily practice in every key area of math instruction Math Minutes is a fun way to improve students problem solving speed crucial for testing situations and an instant assessment tool Holt Biology: The environment ,2003 Canada Map Skills (eBook) David Fletcher,Sharon Bocklage,2003-09-01 This book presents

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