

## WHEEL & TOOLS

Using a jack other than the one supplied with the vehicle, refer page 7-21 for lifting positions.

Wheel, jack, jack handle and wheel wrench are stored inside the boot.

Wheel, jack handle and wheel wrench are on the left side of the boot. Push the wheel upwards to release.

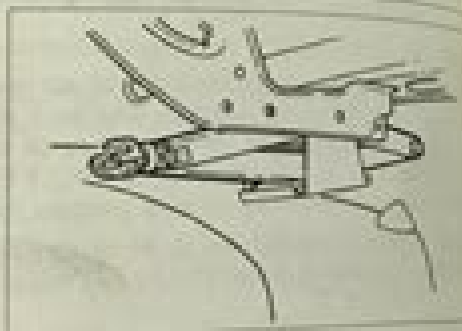
Jack handle, loosen the jack located under the parcel shelf, towards you. The jack can then be pulled out towards you.

When reinstalling the spare wheel, lower strap through buckle on top. Slip wheel wrench through the hole and tighten strap firmly by pulling the strap downwards.

Wheel, wheel wrench, jack handle are stored under the cover in the rear load area.

Open the tailgate and roll the cover forwards.

Wheels - unclip the spare wheel cover then lift cover up. 15" - unscrew the centre



## SPARE WHEEL & TOOLS

### IDENTIFY

Remove the tools from behind the seats:

- wheel wrench
- jack
- jack handle
- spare wheel hoist handle

At the rear of the vehicle insert the spare wheel hoist handle through the round hole near the rear licence plate.

"Jiggle" the hoist handle until the square end of the handle is located in the square cut out in the hoist.

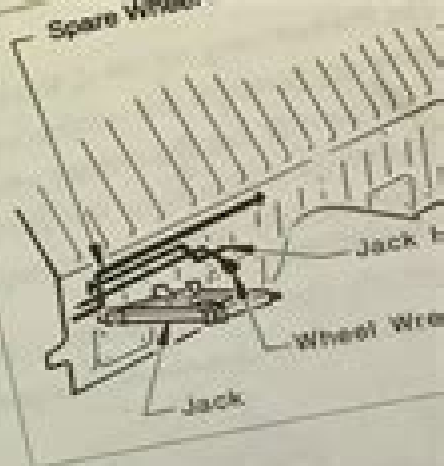
Turn the lever anti-clockwise to fully lower the spare wheel.

Slide the wheel retainer, allowing this to pass through the centre of the wheel.

To replace:

Place the flat tyre under the vehicle, with the tyre

## SPARE WHEEL



# Vp Commodore Owners Manual

**Y Pai**



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web by utilizing the half equivalence point on the titration curve the pka of the acid analyte was determined to be 3.655 furthermore using the stoichiometric ratio between the naoh and the acid analyte 1:1 we could experimentally determine the number of moles of the acid in the solution to be 0.00731 moles making the experimental molar mass

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web goals understand that parts of an acid base titration be able to determine the  $K_a$  or  $K_b$  from pH data associated with the titration of a weak acid or base be able to determine the molar mass of a solid monoprotic acid from titration data be able to

calculate  $K_{a1}$  and  $K_{a2}$  for a polyprotic acid by the end of this lab students should be able to

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web titration experiment in the neutralization of hydrochloric acid by sodium hydroxide the mole ratio of acid to base is 1 : 1  
 $\text{HCl(aq)} + \text{NaOH(aq)} \rightarrow \text{NaCl(aq)} + \text{H}_2\text{O(l)}$   
 one mole of HCl would be fully neutralized by one mole of NaOH

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web specifically an acid base titration can be used to figure out the concentration of an acid or base whether an unknown acid or base is strong or weak  $pK_a$  of an unknown acid or  $pK_b$  of the unknown base key terms titration a process where a solution of known strength is added to a certain volume of a treated sample containing an

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web conclusion this lab evaluated an acid base reaction with the volumetric analysis of a standard solution being NaOH reaction with a 2nd substance  $\text{H}_2\text{C}_2\text{O}_4 + 2\text{NaOH} \rightarrow \text{Na}_2\text{C}_2\text{O}_4 + 2\text{H}_2\text{O}$  to differentiate an acid from a base in a solution a pH scale is used the pH scale ranges from 0 to 14 a pH of 7 is neutral in an acid base reaction

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