

**1a)** When ever a metal is dipped in its own salt solution it has capacity to either loose or gain electrons, due to which it gains some potential which is known as single electrode potential & is represented as 'E'

Nernst equation gives the relation ship between the electrode potential concentration of the species and was derive by a well known scientist Nernst. it was derived by taking well known thermodynamic considerations

Consider a reversible equation



The decrease in free energy is given by Vant hoffs reaction isotherm

$$\Delta G = \Delta G^\circ + RT \ln k \quad 1$$

$\Delta G$  = Gibbs free energy     $\Delta G^\circ$  = Standard gibbs free energy

$R$  = Gas constant     $T$  = Temperature     $k$  = equilibrium constant

Decrease in free energy represents the maximum work done

$$\text{i.e., } -\Delta G = W_{\max}$$

$$\text{i.e., } W_{\max} = nFE$$

where 'n' is the total number of electrons lost during oxidation or gained during reduction

$$\text{therefore } \Delta G = -nFE \quad 2$$

$$\Delta G^\circ = -nFE^\circ \quad 3$$

$k$  =  $\frac{\text{concentration of products}}{\text{concentration of reactants}}$

Substituting 2 & 3 in 1 we get,

$$-nFE = -nFE^\circ + RT \ln \frac{[M]}{[M^{n+}]}$$

Divide the whole equation by  $-nF$  we get,

$$E = E^\circ - \frac{RT}{nF} \ln \frac{[M]}{[M^{n+}]}$$

Converting  $\ln$  to  $\log$  & changing the sign we get

# Vtu Dipmaths 1question Paper

**D. W. Russell, Philip V. Ellery, R. K.  
Williamson**



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