# Specific Heat Capacity Practice Problems

Number:

When 3.0 kg of water is cooled from 80.0°C to 10.0°C, how much heat energy is lost?

$$C_{P_1} = Q/m \cdot \Delta T$$
 = 879,0605  
=  $C_{P_1} \cup (m) \Delta T$  =  $(H_1 \otimes U)(3)(10-80)$ 

How much heat is needed to raise a 0.30 kg piece of aluminum from 30.°C to 150°C?

Calculate the temperature change when:

a) 10.0 kg of water loses 232 kJ of heat.  

$$232 \text{ kJ} = \frac{1600 \text{ J}}{1 \text{ kJ}} = 232000 \text{ J}$$
  
 $-232,000 \text{ J} = (418 \text{ k}) (10.0 \text{ kg}) \text{ kJ}$   
 $-232,000 = 418 \text{ k}0 \text{ AJ}$   
b) 1.96 kJ of heat are added to 500. g of copper.  
 $196 \text{ kJ} \cdot \frac{1000 \text{ J}}{1000 \text{ J}} = 1960 \text{ J}$   
 $196 \text{ kJ} \cdot \frac{1000 \text{ J}}{1000 \text{ J}} = 1960 \text{ J}$ 

196 KJ - 1000 J = 1960 J

4. 2.52 x 104 J of heat are added to 2.0 kg of mercury to reach a final temperature of 130°C. What was the initial temperature of the mercury?

2.5
$$2 \times 10^4 = (138)(2)(130 - T_i)$$
  
2.5 $2 \times 10^4 = 276(130 - T_i)$   
91.3 = 130 - Ti  
-38.7 = -Ti

# **Specific Heat Problems And Answers**

H. Stephen Stoker, Danny V. White

#### **Specific Heat Problems And Answers:**

Problems And Solutions On Thermodynamics And Statistical Mechanics Yung-kuo Lim,1990-02-01 The material for these volumes has been selected from the past twenty years examination questions for graduate students at University of California at Berkeley Columbia University the University of Chicago MIT State University of New York at Buffalo Princeton University and University of Wisconsin Physics with Answers Andrew R. King,Oded Regev,1997-05-28 This book contains 500 problems covering all of introductory physics along with clear step by step solutions to each problem

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