

DCR-SR190E/SR200/SR200C/SR200E/SR290E/ SR300/SR300C/SR300E

RMT-835

SERVICE MANUAL

Ver 1.1 2007.03

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web 1 consider the reaction $\text{i}_2\text{o}_5 \text{ g } 5 \text{ co g } 5 \text{ co}_2 \text{ g } \text{i}_2 \text{ g a } 80 \text{ 0 grams of iodine v oxide i}_2\text{o}_5 \text{ reacts with } 28 \text{ 0 grams of carbon monoxide co determine the mass of iodine i}_2 \text{ which could be produced b if in the above situation only } 0 \text{ 160 moles of iodine i}_2 \text{ was produced i what mass of iodine was produced}$

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$$4\text{Cu} + 2\text{AgNO}_3 \rightarrow 2\text{Cu} + 2\text{Ag} + 2\text{CuNO}_3$$

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$$3\text{CO} + 7\text{H}_2 \rightarrow \text{C}_3\text{H}_8 + 3\text{H}_2\text{O}$$

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web start with 8.3 mole H₂S and multiply it with 2 mole H₂O over 2 mole H₂S from the equation

$$8.3 \text{ mole H}_2\text{S} \times \frac{2 \text{ mole H}_2\text{O}}{2 \text{ mole H}_2\text{S}} = 8.3 \text{ mole H}_2\text{O}$$

8.3 mole H₂S $\times$ 18 g H₂O / 1 mole H₂O = 149.4 g H₂O is your theoretical yield so percent yield = $\frac{137.1 \text{ g H}_2\text{O}}{149.4 \text{ g H}_2\text{O}} \times 100 = 91.77\%$

problems 6 & 8 are similar to problem 5

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$$\text{C}_3\text{H}_6 + 3\text{O}_2 \rightarrow 3\text{CO} + 3\text{H}_2\text{O}$$

if the reaction of 21.3 grams of O₂ produces a 47.3% yield how many grams of CO would be produced

step one figure out how many moles of the limiting reagent you have

$$21.3 \text{ g O}_2 \times \frac{1 \text{ mole O}_2}{32 \text{ g O}_2} = 0.666 \text{ moles O}_2$$

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web percent yield is a measure of the efficiency of the experimental design

$$\text{yield} = \frac{\text{mass of product obtained}}{\text{calculated mass of the product expected}} \times 100$$

in this reaction a yield of 80% is anticipated

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web a write a balanced equation for the reaction

$$\text{Al} + \text{O}_2 \rightarrow \text{Al}_2\text{O}_3$$

4 Al + 3 O₂ $\rightarrow$ 2 Al₂O₃

b determine the theoretical yield of Al₂O₃

$$160.0 \text{ g Al} \times \frac{1 \text{ mol Al}}{27.0 \text{ g Al}} \times \frac{2 \text{ mol Al}_2\text{O}_3}{4 \text{ mol Al}} \times 101.96 \text{ g Al}_2\text{O}_3 = 150.1 \text{ g Al}_2\text{O}_3$$

160.0 g Al $\times$ 1 mol Al / 27.0 g Al $\times$ 2 mol Al₂O₃ / 4 mol Al $\times$ 101.96 g Al₂O₃ = 150.1 g Al₂O₃

302.3 g Al₂O₃ $\times$ 101.96 g Al₂O₃ / 302.3 g Al₂O₃ = 101.96 g Al₂O₃

4 mol Al + 1 mol Al₂O₃

c determine the percent yield

$$\text{yield} = \frac{\text{actual}}{\text{theoretical}} \times 100 = \frac{260.0 \text{ g}}{260.0 \text{ g}} \times 100 = 100\%$$

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