

## Chapter 1

## Problem Solutions

## 1.1

- (a) fcc: 8 corner atoms  $\times 1/8 = 1$  atom  
 6 face atoms  $\times 1/2 = 3$  atoms  
 Total of 4 atoms per unit cell

- (b) bcc: 8 corner atoms  $\times 1/8 = 1$  atom  
 1 enclosed atom  $= 1$  atom  
 Total of 2 atoms per unit cell

- (c) Diamond: 8 corner atoms  $\times 1/8 = 1$  atom  
 6 face atoms  $\times 1/2 = 3$  atoms  
 4 enclosed atoms  $= 4$  atoms  
 Total of 8 atoms per unit cell

## 1.2

- (a) 4 Ga atoms per unit cell

$$\text{Density} = \frac{4}{(5.65 \times 10^{-8})^3} \Rightarrow$$

$$\text{Density of Ga} = 2.22 \times 10^{22} \text{ cm}^{-3}$$

4 As atoms per unit cell, so that

$$\text{Density of As} = 2.22 \times 10^{22} \text{ cm}^{-3}$$

- (b)

8 Ge atoms per unit cell

$$\text{Density} = \frac{8}{(5.65 \times 10^{-8})^3} \Rightarrow$$

$$\text{Density of Ge} = 4.44 \times 10^{22} \text{ cm}^{-3}$$

## 1.3

- (a) Simple cubic lattice;  $a = 2r$

$$\text{Unit cell vol} = a^3 = (2r)^3 = 8r^3$$

$$1 \text{ atom per cell, so atom vol.} = (1) \left( \frac{4\pi r^3}{3} \right)$$

Then

$$\text{Ratio} = \frac{\left( \frac{4\pi r^3}{3} \right)}{8r^3} \times 100\% \Rightarrow \text{Ratio} = 52.4\%$$

- (b) Face-centered cubic lattice

$$d = 4r = a\sqrt{2} \Rightarrow a = \frac{d}{\sqrt{2}} = 2\sqrt{2} r$$

$$\text{Unit cell vol} = a^3 = (2\sqrt{2} r)^3 = 16\sqrt{2} r^3$$

$$4 \text{ atoms per cell, so atom vol.} = 4 \left( \frac{4\pi r^3}{3} \right)$$

Then

$$\text{Ratio} = \frac{4 \left( \frac{4\pi r^3}{3} \right)}{16\sqrt{2} r^3} \times 100\% \Rightarrow \text{Ratio} = 74\%$$

- (c) Body-centered cubic lattice

$$d = 4r = a\sqrt{3} \Rightarrow a = \frac{4}{\sqrt{3}} r$$

$$\text{Unit cell vol.} = a^3 = \left( \frac{4}{\sqrt{3}} r \right)^3$$

$$2 \text{ atoms per cell, so atom vol.} = 2 \left( \frac{4\pi r^3}{3} \right)$$

Then

$$\text{Ratio} = \frac{2 \left( \frac{4\pi r^3}{3} \right)}{\left( \frac{4}{\sqrt{3}} r \right)^3} \times 100\% \Rightarrow \text{Ratio} = 68\%$$

- (d) Diamond lattice

$$\text{Body diagonal} = d = 8r = a\sqrt{3} \Rightarrow a = \frac{8}{\sqrt{3}} r$$

$$\text{Unit cell vol.} = a^3 = \left( \frac{8}{\sqrt{3}} r \right)^3$$

$$8 \text{ atoms per cell, so atom vol.} = 8 \left( \frac{4\pi r^3}{3} \right)$$

Then

$$\text{Ratio} = \frac{8 \left( \frac{4\pi r^3}{3} \right)}{\left( \frac{8}{\sqrt{3}} r \right)^3} \times 100\% \Rightarrow \text{Ratio} = 34\%$$

## 1.4

From Problem 1.3, percent volume of fcc atoms is 74%; Therefore after coffee is ground,

$$\text{Volume} = 0.74 \text{ cm}^3$$

# Solutions Manual 11 Mcgraw Hill Physics

**JS Bruner**



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