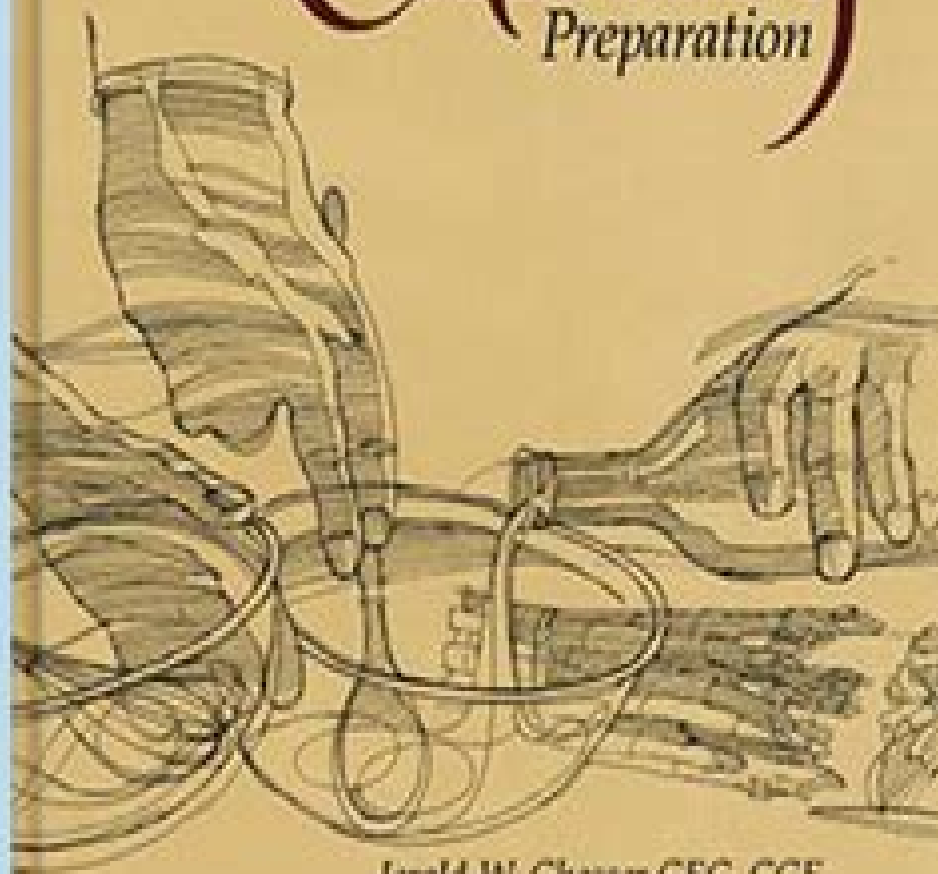


*The Art and
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Jerald W. Chesser CEC, CCE
Educational Institute of Culinary Federation

The Art And Science Of Culinary Preparation

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$$3\text{CO} + 7\text{H}_2\text{C}_3\text{H}_8 \rightarrow 3\text{H}_2\text{O}$$

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 $8.3 \text{ mole } \text{H}_2\text{O} \times 18 \text{ g } \text{H}_2\text{O} = 149.4 \text{ g } \text{H}_2\text{O}$ is your theoretical yield so percent yield $\frac{137.1 \text{ g } \text{H}_2\text{O}}{149.4 \text{ g } \text{H}_2\text{O}} \times 100 = 91.77\%$
problems 6 & 8 are similar to problem 5

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web a write a balanced equation for the reaction $\text{Al} + \text{O}_2 \rightarrow \text{Al}_2\text{O}_3$ b determine the theoretical yield of Al_2O_3 $160.0 \text{ g } \text{Al} \times \frac{1 \text{ mol } \text{Al}}{27 \text{ g } \text{Al}} \times \frac{2 \text{ mol } \text{Al}_2\text{O}_3}{4 \text{ mol } \text{Al}} \times 101.96 \text{ g } \text{Al}_2\text{O}_3/\text{mol} = 150.1 \text{ g } \text{Al}_2\text{O}_3$ c determine the percent yield $\frac{126.0 \text{ g}}{150.1 \text{ g}} \times 100 = 83.9\%$

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