

Chemistry 338 2nd Semester Final Exam Review Sheet

Unit 1 - Matter & Measurement – Chapters 1 & 2

1. A student experimentally determines the density of an object to be 4.2 g/mL. If the accepted value is 5.0 g/mL, calculate the % error.

$$\% \text{ error} = \frac{\text{experimental} - \text{accepted}}{\text{accepted}} \times 100$$

$$\% \text{ error} = \frac{4.2 - 5.0}{5.0} \times 100 = -16\% \text{ error}$$

Unit 2 – Periodic Table – Chapter 5

2.

Element	Symbol	Atomic #	Protons	Neutrons	Electrons	Mass #	Charge
Tungsten ion	¹⁸⁴ ₇₄ W ⁺²	74	74	110	72	184	+2

3. Using the following atomic radius values, predict the atomic radius of potassium (K):
 Li – 145 pm, Na – 180 pm, K - ?, Rb – 235 pm
K – 220 pm

Unit 3 – Reactions/Compounds – Chapters 7 & 8

4. Classify the compound as ionic or covalent. Name the compound
- Fe₂O₃ - iron (III) oxide (ionic - roman numerals for transition metals, change ending to -ide)
 - Cl₂O₈ - dichlorine octaoxide (covalent - use prefixes, change ending to -ide)
5. Classify the compound as ionic or covalent. Write the formula
- Trisulfur Dinitride – S₃N₂ (covalent)
 - Zinc Hydroxide – Zn(OH)₂ (ionic – look up ion charges and criss-cross) Zn⁺² OH⁻¹
 - Chromium (III) nitrate – Cr(NO₃)₃ (ionic – criss-cross) Cr⁺³ NO₃⁻²
6. Write the equation, balance it, and classify the type of reaction:
 Iron (II) sulfate reacts with sodium hydroxide to yield iron(II) hydroxide and sodium sulfate
- $$1 \text{ FeSO}_4 + 2 \text{ NaOH} \rightarrow 1 \text{ Fe(OH)}_2 + 1 \text{ Na}_2\text{SO}_4$$
- double displacement (two ionic compounds reacting)
7. Balance the equation and classify the type of reaction:
- $$1 \text{ Cu} + 2 \text{ AgNO}_3 \rightarrow 1 \text{ Cu(NO}_3)_2 + 2 \text{ Ag}$$
- Single displacement (element and a compound reacting)

Unit 4 – Moles – Chapter 3

8. Calculate the molar mass of Fe₂(SO₄)₃.

$$\begin{aligned} \text{Fe (2)(55.9 g/mol)} &= 111.8 \text{ g/mol} \\ \text{S (3)(32.1 g/mol)} &= 96.3 \text{ g/mol} \\ + \text{O (12)(16.0 g/mol)} &= +192 \text{ g/mol} \\ &= 400.1 \text{ g/mol} \end{aligned}$$

Standard Chemistry Fall Exam Review Sheet

Ruth E. McCall, Cathee M. Tankersley



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hired science teachers different ways in which to support newly hired science teachers and different research approaches that can provide new information about this group of teachers Chapters in the book are written by those who study the status of beginning science teachers mentor new teachers develop induction programs and research the development of new science teachers Newly Hired Teachers of Science is for administrators who have new science teachers in their schools and districts professionals who create science teacher induction programs mentors who work closely with new science teachers educational researchers interested in studying new science teachers and even new science teachers This is a comprehensive discussion about new science teachers that will be a guiding document for years to come

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