

Assessment #3 Review

Name: _____

You must show all work for credit.

Due Date: _____ (No late packets will be accepted)

I. Balance Equations*Balance the equations*

- $\text{Mg} + \underline{2} \text{H}_2\text{O} \rightarrow \text{Mg}(\text{OH})_2 + \text{H}_2$
- $\text{Mg} + \underline{2} \text{NaBr} \rightarrow \text{MgBr}_2 + \underline{2} \text{Na}$
- $\text{C}_3\text{H}_8 + \underline{5} \text{O}_2 \rightarrow \underline{3} \text{CO}_2 + \underline{4} \text{H}_2\text{O}$
- $\underline{4} \text{P} + \underline{5} \text{O}_2 \rightarrow \underline{2} \text{P}_2\text{O}_5$
- $\text{Fe}_2(\text{SO}_4)_3 + \underline{6} \text{KOH} \rightarrow \underline{3} \text{K}_2\text{SO}_4 + \underline{2} \text{Fe}(\text{OH})_3$
- $\text{Fe}_2\text{O}_3 + \underline{3} \text{H}_2 \rightarrow \text{Fe} + \underline{3} \text{H}_2\text{O}$

7. How many moles of hydrogen gas react with 1 mole of Fe_2O_3 in reaction #6? 3

Write the balanced equation for problems # 8 and 9:

8. Hydrogen gas and nitrogen monoxide combine to form water and nitrogen gas



9. When sodium phosphate and calcium chloride are mixed they form calcium phosphate and sodium chloride



10. Using the following balanced equation, prove (or explain) the Law of Conservation of Mass.



2 N on left, 2 N on right, 6 H on left, 6 H on right, 2 on left, 2 on right.

11. Using the balanced equation shown in problem #10, explain what the coefficients represent.

2 N, 6 H, 2 on left, 2 on right.

12. Using the balanced equation shown in problem #10, list the products of the reaction.

2NH₃

13. Using the balanced equation shown in problem #10, list the reactants of the reaction.

N₂, 3H₂

II. Types of Reactions*Directions: Identify the type of reaction.*

S = synthesis

SD = single displacement

C = combustion

D = decomposition

DD = double displacement



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Table of Contents Standard Chemistry Review For Unit 6 Test

1. Understanding the eBook Standard Chemistry Review For Unit 6 Test
 - The Rise of Digital Reading Standard Chemistry Review For Unit 6 Test
 - Advantages of eBooks Over Traditional Books
2. Identifying Standard Chemistry Review For Unit 6 Test
 - Exploring Different Genres
 - Considering Fiction vs. Non-Fiction
 - Determining Your Reading Goals
3. Choosing the Right eBook Platform
 - Popular eBook Platforms
 - Features to Look for in an Standard Chemistry Review For Unit 6 Test
 - User-Friendly Interface
4. Exploring eBook Recommendations from Standard Chemistry Review For Unit 6 Test
 - Personalized Recommendations
 - Standard Chemistry Review For Unit 6 Test User Reviews and Ratings
 - Standard Chemistry Review For Unit 6 Test and Bestseller Lists
5. Accessing Standard Chemistry Review For Unit 6 Test Free and Paid eBooks
 - Standard Chemistry Review For Unit 6 Test Public Domain eBooks
 - Standard Chemistry Review For Unit 6 Test eBook Subscription Services
 - Standard Chemistry Review For Unit 6 Test Budget-Friendly Options

6. Navigating Standard Chemistry Review For Unit 6 Test eBook Formats
 - ePub, PDF, MOBI, and More
 - Standard Chemistry Review For Unit 6 Test Compatibility with Devices
 - Standard Chemistry Review For Unit 6 Test Enhanced eBook Features
7. Enhancing Your Reading Experience
 - Adjustable Fonts and Text Sizes of Standard Chemistry Review For Unit 6 Test
 - Highlighting and Note-Taking Standard Chemistry Review For Unit 6 Test
 - Interactive Elements Standard Chemistry Review For Unit 6 Test
8. Staying Engaged with Standard Chemistry Review For Unit 6 Test
 - Joining Online Reading Communities
 - Participating in Virtual Book Clubs
 - Following Authors and Publishers Standard Chemistry Review For Unit 6 Test
9. Balancing eBooks and Physical Books Standard Chemistry Review For Unit 6 Test
 - Benefits of a Digital Library
 - Creating a Diverse Reading Collection Standard Chemistry Review For Unit 6 Test
10. Overcoming Reading Challenges
 - Dealing with Digital Eye Strain
 - Minimizing Distractions
 - Managing Screen Time
11. Cultivating a Reading Routine Standard Chemistry Review For Unit 6 Test
 - Setting Reading Goals Standard Chemistry Review For Unit 6 Test
 - Carving Out Dedicated Reading Time
12. Sourcing Reliable Information of Standard Chemistry Review For Unit 6 Test
 - Fact-Checking eBook Content of Standard Chemistry Review For Unit 6 Test
 - Distinguishing Credible Sources
13. Promoting Lifelong Learning
 - Utilizing eBooks for Skill Development
 - Exploring Educational eBooks
14. Embracing eBook Trends
 - Integration of Multimedia Elements

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