Heat of Formation Worksheet

Use a standard enthalpies of formation table to determine the change in enthalpy for each of these reactions.

a) $NaOH(s) + HCl(g) \longrightarrow NaCl(s) + H_2O(g)$

$$[(-411.0) + (-241.8)] - [(-426.7) + (-92.3)] = -133.8 \text{ kJ}$$

b) 2 CO(g) + O₂(g) ---> 2 CO₂(g)

c) CH₄(g) + 2 O₂(g) ---> CO₂(g) + 2 H₂O(l)

$$[(-393.5) + 2(-285.8)] - [(-74)] = -891.1 \text{ kJ}$$

d) $2 H_2S(g) + 3 O_2(g) \longrightarrow 2 H_2O(1) + 2 SO_2(g)$

$$[2(-285.8) + 2(-296.1)] - [2(-20.1)] = -1123.6 \text{ kJ}$$

e) 2 NO(g) + O₂(g) ---> 2 NO₂(g)

$$[2(+33.9)] - [2(+90.4)] = -113 \text{ kJ}$$

Compound	AHe (kJ/mol)	Compound	□Hr (kJ/mol)
CH _r (g)	-74.8	HCl(g)	-92.3
CO ₂ (g)	-393.5	H ₂ O(g)	-241.8
NaCl(s)	-411.0	SO ₂ (g)	-296.1
H ₂ O(I)	-285.8	NH ₆ Cl(s)	-315.4
H ₂ S(g)	-20.1	NO(g)	+90.4
H ₂ SO ₄ (I)	-811.3	NO ₂ (g)	+33.9
MgSO ₄ (s)	-1278.2	SnCl ₄ (I)	-545.2
MmO(s)	-384.9	SnO(s)	-286.2
MmO ₂ (s)	-519.7	SarO ₂ (x)	-580.7
NaCl(s)	-411.0	SO ₂ (g)	-296.1
NaF(s)	-569.0	SO ₂ (g)	-395.2
NaOH(s)	-426.7	ZmO(s)	-348.0
NH ₂ (g)	-46.2	ZmS(s)	-202.9

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