

Heat of Formation Worksheet

Use a standard enthalpies of formation table to determine the change in enthalpy for each of these reactions.



$$[(-411.0) + (-241.8)] - [(-426.7) + (-92.3)] = -133.8 \text{ kJ}$$



$$[2(-393.5)] - [2(-110.5)] = -566 \text{ kJ}$$



$$[(-393.5) + 2(-285.8)] - [(-74)] = -891.1 \text{ kJ}$$



$$[2(-285.8) + 2(-296.1)] - [2(-20.1)] = -1123.6 \text{ kJ}$$



$$[2(+33.9)] - [2(+90.4)] = -113 \text{ kJ}$$

| Compound | ΔH_f (kJ/mol) | Compound | ΔH_f (kJ/mol) |
|-----------------------------------|-----------------------|---------------------------|-----------------------|
| $\text{CH}_4\text{(g)}$ | -74.8 | HCl(g) | -92.3 |
| $\text{CO}_2\text{(g)}$ | -393.5 | $\text{H}_2\text{O(g)}$ | -241.8 |
| NaCl(s) | -411.0 | $\text{SO}_2\text{(g)}$ | -296.1 |
| $\text{H}_2\text{O(l)}$ | -285.8 | $\text{NH}_4\text{Cl(s)}$ | -315.4 |
| $\text{H}_2\text{S(g)}$ | -20.1 | NO(g) | +90.4 |
| $\text{H}_2\text{SO}_4\text{(l)}$ | -811.3 | $\text{NO}_2\text{(g)}$ | +33.9 |
| $\text{MgSO}_4\text{(s)}$ | -1278.2 | $\text{SnCl}_4\text{(l)}$ | -545.2 |
| MnO(s) | -384.9 | SnO(s) | -286.2 |
| $\text{MnO}_2\text{(s)}$ | -519.7 | $\text{SnO}_2\text{(s)}$ | -580.7 |
| NaCl(s) | -411.0 | $\text{SO}_3\text{(g)}$ | -296.1 |
| NaF(s) | -569.0 | $\text{SO}_3\text{(g)}$ | -395.2 |
| NaOH(s) | -426.7 | ZnO(s) | -348.0 |
| $\text{NH}_3\text{(g)}$ | -46.2 | ZnS(s) | -202.9 |

Standard Heat Of Formation Answer Key

Disha Experts

A red circular graphic with a gradient, appearing as a semi-circle or a partial circle, located to the right of the 'Disha Experts' text.

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