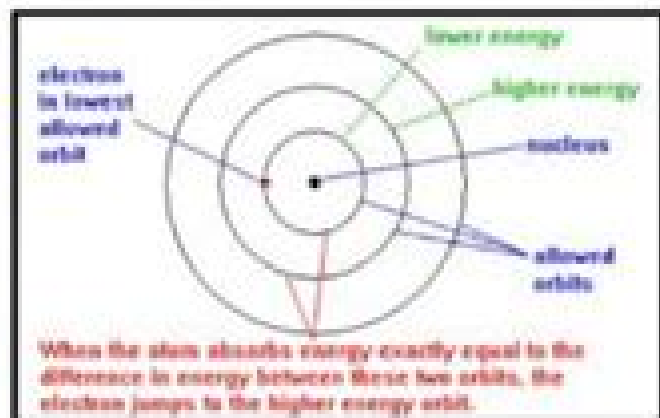


# The Bohr Model

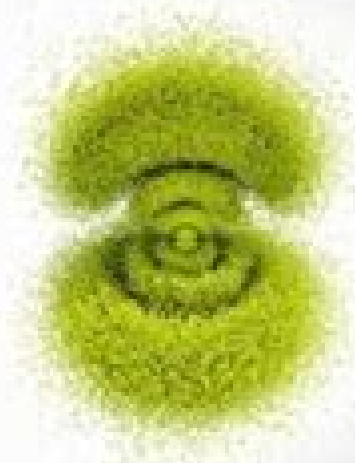
A one-dimensional model, one quantum number ( $n$ ) to describe the distribution of electrons in the atom.



3



what  
atoms  
really  
look like?



4



## Quantum electron configuration

**Angular ( $l$ ):** describes the shape of the orbital.

- Orbitals have shapes that are best described as spherical ( $l = 0$ ), polar ( $l = 1$ ), or cloverleaf ( $l = 2$ ).
- They can even take on more complex shapes as the value of the angular quantum number becomes larger.

Angular Quantum Number, $l$	Name of Orbital	Shape
0	s	spherical
1	p	dumbbell
2	d	complex/cloverleaf
3	f	even more complex/multiple lobes

## Quantum electron configuration

**Magnetic ( $m_l$ ):** describes the orientation in space

- There is only one way in which a sphere ( $l = 0$ ) can be oriented in space.
- $l$  is called the magnetic quantum number because the effect of different orientations of orbitals was first observed in the presence of a magnetic field.



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