

Chemical Equilibrium

Multiple Choice Practice

1. Chemical reactions are in a state of dynamic equilibrium only when
- A. the rate of the forward reaction equals that of the backward reaction.
 - B. the concentration of reactants and products are equal
 - C. the activation energies of the forward and backward reactions are equal
 - D. the reaction involves no enthalpy change.
2. Under the conditions used industrially, ethene and steam react as follows:



Which set of conditions would give the best yield of ethanol at equilibrium?

- A. high temperature, low pressure
 - B. high temperature, high pressure
 - C. low temperature, high pressure
 - D. low temperature, low pressure.
3. $\text{N}_2(\text{g}) + 3\text{H}_2(\text{g}) \rightleftharpoons 2\text{NH}_3(\text{g})$
- In the Haber process, illustrated by the equation, the reaction never reaches equilibrium because
- A. the ammonia is constantly removed
 - B. an iron catalyst is used
 - C. high pressure is used
 - D. the temperature is held at above 500°C .

4. The equation refers to the preparation of methanol from synthesis gas.
- $$\text{CO}(\text{g}) + 2\text{H}_2(\text{g}) \rightleftharpoons \text{CH}_3\text{OH}(\text{g})$$
- $$\Delta H = -91 \text{ kJ mol}^{-1}$$

The formation of methanol is favoured by

- A. high pressure and low temperature
 - B. high pressure and high temperature
 - C. low pressure and low temperature
 - D. low pressure and high temperature.
5. $\text{ICl}(\text{l}) + \text{Cl}_2(\text{g}) \rightleftharpoons \text{ICl}_3(\text{s})$
- $$\Delta H = -46 \text{ kJ mol}^{-1}$$

Which of the following sets of changes will cause the greatest increase in the proportion of solid in the above equilibrium?

	Temperature	Pressure
A	decrease	decrease
B	decrease	increase
C	increase	decrease
D	increase	increase

6. When a reversible chemical reaction is at equilibrium,
- A. the concentration of reactants and products remain equal
 - B. the forward reaction is unable to continue
 - C. the concentration of reactants and products remain constant
 - D. the forward and reverse reactions proceed at different rates.

Static Equilibrium Multiple Choice

John Bird



Static Equilibrium Multiple Choice:

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Mechanical Engineering Principles John Bird, Carl Ross, 2014-11-27 A student friendly introduction to core engineering topics This book introduces mechanical principles and technology through examples and applications enabling students to develop a sound understanding of both engineering principles and their use in practice These theoretical concepts are supported by 400 fully worked problems 700 further problems with answers and 300 multiple choice questions all of which add up to give the reader a firm grounding on each topic The new edition is up to date with the latest BTEC National specifications and can also be used on undergraduate courses in mechanical civil structural aeronautical and marine engineering together with naval architecture A further chapter has been added on revisionary mathematics since progress in engineering studies is not possible without some basic mathematics knowledge Further worked problems have also been added throughout the text New chapter on revisionary mathematics Student friendly approach with numerous worked problems multiple choice and short answer questions exercises revision tests and nearly 400 diagrams Supported with free online material for students and lecturers Readers will also be able to access the free companion website where

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