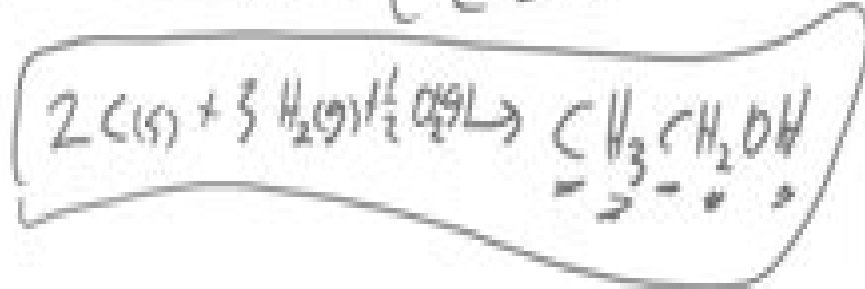


Calculation of $\Delta_f H^\circ$ Self-test

Which equation below defines the standard molar enthalpy of formation of ethanol, $\text{CH}_3\text{CH}_2\text{OH}$?

$P = 1 \text{ bar}$
 $T = 298 \text{ K}$
 $C = 1 \text{ M}$

- a) $\text{C}_2\text{H}_6(\text{g}) + \frac{1}{2} \text{O}_2(\text{g}) \rightarrow \text{CH}_3\text{CH}_2\text{OH}(\text{l})$
- b) $2 \text{C}(\text{s}) + 3 \text{H}_2(\text{g}) + \frac{1}{2} \text{O}_2(\text{g}) \rightarrow \text{CH}_3\text{CH}_2\text{OH}(\text{l})$
- c) $\text{CO}(\text{g}) + \text{CH}_4(\text{g}) + \text{H}_2(\text{g}) \rightarrow \text{CH}_3\text{CH}_2\text{OH}(\text{g})$
- d) $\text{H}_2\text{O}(\text{g}) + 2 \text{C}(\text{s}) + 2 \text{H}_2(\text{g}) \rightarrow \text{CH}_3\text{CH}_2\text{OH}(\text{l})$
- e) $\text{CH}_3\text{CH}_2\text{OH}(\text{l}) + 3 \text{O}_2(\text{g}) \rightarrow 2 \text{CO}_2(\text{g}) + 3 \text{H}_2\text{O}(\text{g})$



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Minjie Lin



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