

AP Chemistry: Chapter 14 Practice Test

10. Given the following acids and K_a values:



$$1 \times 10^7$$

$$1.76 \times 10^{-5}$$

$$4.93 \times 10^{-10}$$

$$3.53 \times 10^{-4}$$

which shows the conjugate bases listed by increasing strength?



15. Given that the K_a for HOCl is 3.5×10^{-8} , calculate the K value for the reaction of HOCl with OH^- .

[A] 3.5×10^{-22}

[B] 3.5×10^6

[C] 2.9×10^{-7}

[D] 3.5×10^{-8}

[E] none of these

32. Acetic acid ($\text{HC}_2\text{H}_3\text{O}_2$) is a weak acid ($K_a = 1.8 \times 10^{-5}$). Calculate the pH of a 17.6 M $\text{HC}_2\text{H}_3\text{O}_2$ solution.

[A] 4.3

[B] 6.4

[C] 1.7

[D] 0.97

[E] 7.4

52. What concentration of acetic acid ($K_a = 1.80 \times 10^{-5}$) has the same pH as that of 5.00×10^{-3} M HCl ?

[A] 5.00×10^{-3} M

[B] 3.60×10^{-3} M

[C] 1.39 M

[D] 5.00 M

[E] none of these

Consider the reaction $\text{HOCl} + \text{F}^- \rightleftharpoons \text{HF} + \text{OCl}^-$

Standardized Test Practice Chapter 14 Chemistry Answers

Michael Brown



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