

1. What is the percent by mass of oxygen in MgO?

- | | | | |
|-------------------------|-----|-------------------------|-----|
| <input type="radio"/> A | 50% | <input type="radio"/> B | 60% |
| <input type="radio"/> C | 20% | <input type="radio"/> D | 40% |

2. How many molecules are in 2.5 mol of NaCl?

- | | | | |
|-------------------------|----------------|-------------------------|---------------------------------|
| <input type="radio"/> A | 4.15 molecules | <input type="radio"/> B | 1.51×10^{24} molecules |
| <input type="radio"/> C | 1.50 molecules | <input type="radio"/> D | 1.50×10^{24} molecules |

3. A student calculated the number of moles of methane (CH₄) present in a 100g sample of a gaseous mixture. If the student determined that there's 0.6 moles of methane present in the sample, what is the percent composition of methane in the sample?

- | | | | |
|-------------------------|------|-------------------------|-------|
| <input type="radio"/> A | 9.6% | <input type="radio"/> B | .096% |
| <input type="radio"/> C | .9% | <input type="radio"/> D | 90% |

4. A 15.67 g sample of a hydrate of magnesium carbonate was heated to drive off the water. The mass was reduced to 7.58 g. What is the formula of the hydrate?

- | | | | |
|-------------------------|---------------------------------------|-------------------------|---------------------------------------|
| <input type="radio"/> A | MgCO ₃ · 5H ₂ O | <input type="radio"/> B | MgCO ₃ · 4H ₂ O |
| <input type="radio"/> C | MgCO ₃ · 6H ₂ O | <input type="radio"/> D | MgCO ₃ · 1H ₂ O |

5. What is the empirical formula of a compound that is 81.82% carbon and 18.18% hydrogen?

- | | | | |
|-------------------------|--------------------------------|-------------------------|-------------------------------|
| <input type="radio"/> A | C ₄ H ₁₀ | <input type="radio"/> B | C ₂ H ₂ |
| <input type="radio"/> C | CH ₄ | <input type="radio"/> D | C ₃ H ₈ |

6. $2 \text{KClO}_3 \rightarrow 2 \text{KCl} + 3 \text{O}_2$

How many moles of oxygen are produced when 6.7 moles of KClO₃ decompose completely?

Star Test 10th Grade Chemistry

Padhraic Smyth



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