

1. What explains the very high melting and boiling point of water

- | | |
|--|--|
| <input type="radio"/> A Strong dipole-dipole bonds between water molecules | <input type="radio"/> B Dispersion forces which are present in all molecules |
| <input type="radio"/> C Strong hydrogen bonds between water molecules | <input type="radio"/> D Asymmetrical shape of the polar bonds. |

2. Does HCl have hydrogen bonding?

- | | |
|----------------------------|-----------------------------|
| <input type="radio"/> A no | <input type="radio"/> B yes |
|----------------------------|-----------------------------|

3. Does CH₄ have hydrogen bonding?

- | | |
|-----------------------------|----------------------------|
| <input type="radio"/> A yes | <input type="radio"/> B no |
|-----------------------------|----------------------------|

4. All molecules have London forces between them, but dipole-dipole and hydrogen bonding are so much stronger that when they are present we can ignore London forces. Which of these has ONLY London forces?

- | | |
|---|--|
| <input type="radio"/> A I ₂ | <input type="radio"/> B OCl ₂ |
| <input type="radio"/> C NH ₃ | <input type="radio"/> D SH ₂ |

5. Which substance would have the weakest intermolecular forces of attraction?

- | | |
|---|--|
| <input type="radio"/> A CH ₄ | <input type="radio"/> B H ₂ O |
| <input type="radio"/> C NaCl | <input type="radio"/> D MgF ₂ |

6. Intermolecular forces for: NH₃



- | | |
|---------------------------------------|--|
| <input type="radio"/> A Dipole dipole | <input type="radio"/> B Dispersion Force |
|---------------------------------------|--|

States Of Matter And Intermolecular Forces Test

K Morrison



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