


Section A: Principles of Chemistry

A1 States of Matter

No.	Answers	Further explanations
1	C	
2	D	$\text{NH}_3(\text{g}) + \text{HCl}(\text{g}) \longrightarrow \text{NH}_4\text{Cl}(\text{s})$ <p>The concentrated ammonia solution gives off ammonia gas and the concentrated hydrochloric acid gives off hydrogen chloride gas, shown by the state symbols (g). The particles of both gases diffuse along the tube, collide and react to form a white ring composed of solid white ammonium chloride particles, shown by the state symbol (s).</p>
3	A	<p>Ammonia molecules diffuse faster than hydrogen chloride molecules. Ammonia molecules are smaller and lighter than hydrogen chloride molecules, so they diffuse faster along the tube, and the particles of the two gases collide and react closer to the source of the hydrogen chloride.</p>
4	C	
5	A	<p>Increased in length and become rigid</p> <p>The paw-paw strip is composed of living cells. Each cell is surrounded by a differentially permeable membrane and contains cytoplasm containing about 80% water. By osmosis, water enters each cell causing it to swell slightly. This causes the strip to increase in length and become rigid.</p>
6	B	<p>Moved down</p> <p>The 20% sucrose solution inside the thistle funnel contains more water than the 40% sucrose solution in the beaker. By osmosis, water moves out of the solution in the thistle funnel through the differentially permeable membrane into the solution in the beaker. This decreases the volume of the solution in the thistle funnel which causes the meniscus to move down.</p>

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Wasserman, Peter V. Minorsky**



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Chemistry: The Central Science Theodore L. Brown, H. Eugene LeMay Jr., Bruce E. Bursten, Catherine Murphy, Patrick Woodward, Steven Langford, Darius Sagatys, Adrian George, 2013-10-04 If you think you know the Brown LeMay Bursten Chemistry text think again In response to market request we have created the third Australian edition of the US bestseller Chemistry The Central Science An extensive revision has taken this text to new heights Triple checked for scientific accuracy and consistency this edition is a more seamless and cohesive product yet retains the clarity innovative pedagogy functional problem solving and visuals of the previous version All artwork and images are now consistent in quality across the entire text And with a more traditional and logical organisation of the Organic Chemistry content this comprehensive text is the source of all the information and practice problems students are likely to need for conceptual understanding development of problem solving skills reference and test preparation

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Andrew Hunt, Graham Curtis, Graham Hill, 2019-07-15 Develop and assess your students knowledge and skills throughout A level with worked examples practical assessment guidance and differentiated end of topic questions in this updated all in one textbook for Years 1 and 2 Combining everything your students need to know for the Pearson Edexcel A level Chemistry specification this revised textbook will Identify the level of your students understanding with diagnostic questions and a summary of prior knowledge at the start of the Student Book Provide support for all 16 required practicals with various activities and questions along with a Practical chapter covering procedural understanding and key ideas related to measurement Improve mathematical skills with plenty of worked examples including notes on methods to help explain the strategies for solving each type of problem Offer plenty of practice with Test yourself questions to help students assess their understanding and measure progress Encourage further reading and study with short passages of extension material Develop understanding with free online access to Test yourself answers and an extended glossary

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Table of Contents States Of Matter Answer Key Chemistry Pearson

1. Understanding the eBook States Of Matter Answer Key Chemistry Pearson
 - The Rise of Digital Reading States Of Matter Answer Key Chemistry Pearson
 - Advantages of eBooks Over Traditional Books
2. Identifying States Of Matter Answer Key Chemistry Pearson
 - Exploring Different Genres
 - Considering Fiction vs. Non-Fiction
 - Determining Your Reading Goals
3. Choosing the Right eBook Platform
 - Popular eBook Platforms
 - Features to Look for in an States Of Matter Answer Key Chemistry Pearson
 - User-Friendly Interface
4. Exploring eBook Recommendations from States Of Matter Answer Key Chemistry Pearson

- Personalized Recommendations
- States Of Matter Answer Key Chemistry Pearson User Reviews and Ratings
- States Of Matter Answer Key Chemistry Pearson and Bestseller Lists
- 5. Accessing States Of Matter Answer Key Chemistry Pearson Free and Paid eBooks
 - States Of Matter Answer Key Chemistry Pearson Public Domain eBooks
 - States Of Matter Answer Key Chemistry Pearson eBook Subscription Services
 - States Of Matter Answer Key Chemistry Pearson Budget-Friendly Options
- 6. Navigating States Of Matter Answer Key Chemistry Pearson eBook Formats
 - ePub, PDF, MOBI, and More
 - States Of Matter Answer Key Chemistry Pearson Compatibility with Devices
 - States Of Matter Answer Key Chemistry Pearson Enhanced eBook Features
- 7. Enhancing Your Reading Experience
 - Adjustable Fonts and Text Sizes of States Of Matter Answer Key Chemistry Pearson
 - Highlighting and Note-Taking States Of Matter Answer Key Chemistry Pearson
 - Interactive Elements States Of Matter Answer Key Chemistry Pearson
- 8. Staying Engaged with States Of Matter Answer Key Chemistry Pearson
 - Joining Online Reading Communities
 - Participating in Virtual Book Clubs
 - Following Authors and Publishers States Of Matter Answer Key Chemistry Pearson
- 9. Balancing eBooks and Physical Books States Of Matter Answer Key Chemistry Pearson
 - Benefits of a Digital Library
 - Creating a Diverse Reading Collection States Of Matter Answer Key Chemistry Pearson
- 10. Overcoming Reading Challenges
 - Dealing with Digital Eye Strain
 - Minimizing Distractions
 - Managing Screen Time
- 11. Cultivating a Reading Routine States Of Matter Answer Key Chemistry Pearson
 - Setting Reading Goals States Of Matter Answer Key Chemistry Pearson
 - Carving Out Dedicated Reading Time
- 12. Sourcing Reliable Information of States Of Matter Answer Key Chemistry Pearson

- Fact-Checking eBook Content of States Of Matter Answer Key Chemistry Pearson
- Distinguishing Credible Sources

13. Promoting Lifelong Learning

- Utilizing eBooks for Skill Development
- Exploring Educational eBooks

14. Embracing eBook Trends

- Integration of Multimedia Elements
- Interactive and Gamified eBooks

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