

## ▶ 실전 연습 해답

### 제2장

#### Self-Check Exercise 2.1

$$357 = 3.57 \times 10^2$$

$$0.0055 = 5.5 \times 10^{-4}$$

#### Self-Check Exercise 2.2

- Three significant figures. The leading zeros (to the left of the 1) do not count, but the trailing zeros do.
- Five significant figures. The one captive zero and the two trailing zeros all count.
- This is an exact number obtained by counting the cars. It has an unlimited number of significant figures.

#### Self-Check Exercise 2.3

- $12.6 \times 0.53 = 6.678 = 6.7$   
 $\begin{array}{r} 12.6 \\ \times 0.53 \\ \hline 378 \\ 630 \\ \hline 6.678 \end{array}$
- $12.6 \times 0.53 = 6.7$   
 $\begin{array}{r} 12.6 \\ \times 0.53 \\ \hline 378 \\ 630 \\ \hline 6.678 \end{array}$
- $\begin{array}{r} 25.36 \\ -4.15 \\ \hline 21.21 \end{array}$      $\begin{array}{r} 21.21 \\ \times 2.317 \\ \hline 9.11408 = 9.114 \end{array}$

#### Self-Check Exercise 2.4

$$0.750 \text{ qt} \times \frac{1.06 \text{ qt}}{1 \text{ qt}} = 0.795 \text{ qt}$$

#### Self-Check Exercise 2.5

$$225 \frac{\text{mi}}{\text{h}} \times \frac{1760 \text{ yd}}{1 \text{ mi}} \times \frac{1 \text{ yd}}{1.094 \text{ yd}} \times \frac{1 \text{ km}}{1000 \text{ yd}} = 362 \frac{\text{km}}{\text{h}}$$

#### Self-Check Exercise 2.6

The best way to solve this problem is to convert 172 K to Celsius degrees. To do this we will use the formula  $T_{\text{C}} = T_{\text{K}} - 273$ .

In this case

$$T_{\text{C}} = T_{\text{K}} - 273 = 172 - 273 = -101$$

So 172 K = -101 °C, which is a lower temperature than -75 °C. Thus 172 K is colder than -75 °C.

#### Self-Check Exercise 2.7

The problem is  $41 \text{ °C} = ? \text{ °F}$ .

Using the formula

$$T_{\text{F}} = 1.80(T_{\text{C}}) + 32$$

we have

$$T_{\text{F}} = 1.80(41) + 32 = 74 + 32 = 106$$

That is,  $41 \text{ °C} = 106 \text{ °F}$ .

#### Self-Check Exercise 2.8

This problem can be stated as  $239 \text{ °F} = ? \text{ °C}$ .

Using the formula

$$T_{\text{C}} = \frac{T_{\text{F}} - 32}{1.80}$$

we have in this case

$$T_{\text{C}} = ? \text{ °C} = \frac{239 - 32}{1.80} = \frac{207}{1.80} = 115$$

That is,  $239 \text{ °F} = 115 \text{ °C}$ .

#### Self-Check Exercise 2.9

We obtain the density of the cleaner by dividing its mass by its volume.

$$\text{Density} = \frac{\text{mass}}{\text{volume}} = \frac{28.1 \text{ g}}{35.8 \text{ mL}} = 0.785 \text{ g/mL}$$

This density identifies the liquid as isopropyl alcohol.

### 제3장

#### Self-Check Exercise 3.1

Items (a) and (c) are physical properties. When the solid gallium melts, it forms liquid gallium. There is no change in composition. Items (b) and (d) reflect the ability to change composition and are thus chemical properties. Statement (b) means that platinum does not react with oxygen to form some new substance. Statement (d) means that copper does react in the air to form a new substance, which is green.

#### Self-Check Exercise 3.2

- Milk turns sour because new substances are formed. This is a chemical change.
- Melting the wax is a physical change (a change of state). When the wax burns, new substances are formed. This is a chemical change.

#### Self-Check Exercise 3.3

- Maple syrup is a homogeneous mixture of sugar and other substances dispersed uniformly in water.
- Helium and oxygen form a homogeneous mixture.
- Oil and vinegar salad dressing is a heterogeneous mixture. (Note the two distinct layers the next time you look at a bottle of dressing.)
- Common salt is a pure substance (sodium chloride), so it always has the same composition. (Note that other substances such as iodine are often added to commercial preparations of table salt, which is mostly sodium chloride. Thus commercial table salt is a homogeneous mixture.)

### 제4장

#### Self-Check Exercise 4.1

- $\text{P}_2\text{O}_{10}$     b.  $\text{UF}_6$     c.  $\text{AsCl}_3$

#### Self-Check Exercise 4.2

In the symbol  $^{90}_{38}\text{Sr}$ , the number 38 is the atomic number, which represents the number of protons in the nucleus of a strontium atom. Because the atom is neutral overall, it must also have 38 electrons. The number 90 (the mass number) represents the number of protons plus the number of neutrons. Thus the number of neutrons is  $A - Z = 90 - 38 = 52$ .

#### Self-Check Exercise 4.3

The atom  $^{201}_{80}\text{Hg}$  has 80 protons, 80 electrons, and  $201 - 80 = 121$  neutrons.

#### Self-Check Exercise 4.4

The atomic number for phosphorus is 15 and the mass number is  $15 + 17 = 32$ . Thus the symbol for the atom is  $^{32}_{15}\text{P}$ .

# Zumdahl Chemistry Answer Key

**Iris Stovall, Donald J. DeCoste**



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## **Table of Contents Zumdahl Chemistry Answer Key**

1. Understanding the eBook Zumdahl Chemistry Answer Key
  - The Rise of Digital Reading Zumdahl Chemistry Answer Key
  - Advantages of eBooks Over Traditional Books
2. Identifying Zumdahl Chemistry Answer Key
  - Exploring Different Genres
  - Considering Fiction vs. Non-Fiction
  - Determining Your Reading Goals
3. Choosing the Right eBook Platform
  - Popular eBook Platforms
  - Features to Look for in an Zumdahl Chemistry Answer Key
  - User-Friendly Interface
4. Exploring eBook Recommendations from Zumdahl Chemistry Answer Key
  - Personalized Recommendations
  - Zumdahl Chemistry Answer Key User Reviews and Ratings
  - Zumdahl Chemistry Answer Key and Bestseller Lists

5. Accessing Zumdahl Chemistry Answer Key Free and Paid eBooks
  - Zumdahl Chemistry Answer Key Public Domain eBooks
  - Zumdahl Chemistry Answer Key eBook Subscription Services
  - Zumdahl Chemistry Answer Key Budget-Friendly Options
6. Navigating Zumdahl Chemistry Answer Key eBook Formats
  - ePub, PDF, MOBI, and More
  - Zumdahl Chemistry Answer Key Compatibility with Devices
  - Zumdahl Chemistry Answer Key Enhanced eBook Features
7. Enhancing Your Reading Experience
  - Adjustable Fonts and Text Sizes of Zumdahl Chemistry Answer Key
  - Highlighting and Note-Taking Zumdahl Chemistry Answer Key
  - Interactive Elements Zumdahl Chemistry Answer Key
8. Staying Engaged with Zumdahl Chemistry Answer Key
  - Joining Online Reading Communities
  - Participating in Virtual Book Clubs
  - Following Authors and Publishers Zumdahl Chemistry Answer Key
9. Balancing eBooks and Physical Books Zumdahl Chemistry Answer Key
  - Benefits of a Digital Library
  - Creating a Diverse Reading Collection Zumdahl Chemistry Answer Key
10. Overcoming Reading Challenges
  - Dealing with Digital Eye Strain
  - Minimizing Distractions
  - Managing Screen Time
11. Cultivating a Reading Routine Zumdahl Chemistry Answer Key
  - Setting Reading Goals Zumdahl Chemistry Answer Key
  - Carving Out Dedicated Reading Time
12. Sourcing Reliable Information of Zumdahl Chemistry Answer Key
  - Fact-Checking eBook Content of Zumdahl Chemistry Answer Key
  - Distinguishing Credible Sources
13. Promoting Lifelong Learning

- Utilizing eBooks for Skill Development
  - Exploring Educational eBooks
14. Embracing eBook Trends
- Integration of Multimedia Elements
  - Interactive and Gamified eBooks

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