- Calculate the mass of Al₂S₃ produced when 100 mL of 0.50 M AlCl₃ is added to 100 mL of 0.50 M Na₂S.
- You are given the equation: AgBr + 2S₂O₃² → Ag(S₂O₃)₂³ + Br². What mass of AgBr can be dissolved by 750, mL of 0.300 M Na₂S₂O₃?
- Given the following chemical equation, determine the theoretical yield of B₂H₆ if exactly 100.0 g of LiAlH₄ was allowed to react with 225 g of BF₃.

Section 4.8

- 39. What volume of 0.1379 M HCl is required to neutralize 10.0 mL of 0.2789 M NaOH solution?
- How many mL of 1.50 M NaOH is required to neutralize 275 mL of 0.5 M H₂SO₆?
- 41. What is the molarity of a solution of HCl if it requires 29.31 mL of a 0.0923 M NaOH solution to reach a phenolphthalein endpoint for the titration of a 10.0 mL aliquot of the HCl solution?
- Given the following two equations, determine the number of grams of manganese dioxide required to prepare enough chlorine gas to produce 25.0 g of potassium hypochlorite.

$$MnO_2 + 4HCI \rightarrow MnCI_2 + CI_2 + 2H_2O$$

 $CI_2 + 2KOH \rightarrow KCI + KCIO + H_2O$

- 43. A titration is done using 0.1302 M NaOH to determine the molar mass of an acid. The acid contains one acidic hydrogen per molecule. If 1.863 g of the acid require 70.11 mL of the NaOH solution, what is the molar mass of the acid?
- A 2.000-g sample of silver alloy was dissolved in nitric acid and then precipitated as AgBr. After drying, the sample of silver bromide weighed 2.000 g. Calculate the percentage of silver in the alloy.
- 45. Complete and balance each acid-base equation (assume complete neutralization):

a.
$$H_2SO_4 + N_0OH \rightarrow$$

b. $H_3PO_4 + Mg(OH)_2 \rightarrow$
c. $H_2SO_3 + N_0OH \rightarrow$
d. $HC_2H_3O_2 + Ba(OH)_2 \rightarrow$

- What volume of 0.2 M NaOH is required to neutralize 50 mL of 0.1 M H₂SO₃?
- How many mL of 2.3 M HNO₃ is needed to neutralize 0.92 L of 0.5 M KOH?
- 48. The neutralization of a 25.00-mL sample of an unknown base requires 18.34 mL of 0.100 M HCl. Assuming that the acid-base stoichiometry is 1:1, what is the concentration of the unknown base?
- 49. Pennies made after 1982 contain about 97% zinc by mass. A student wants to prove this by filing the copper outside of a penny until he sees zinc and then putting the penny in a 1.00 M HCl solution. The zinc will be oxidized, and the H* (from HCl) will be reduced:

$$Zn + 2H^* \rightarrow Zn^{2*} + H_2$$

If the entire penny has a mass of 2.80 grams, how many mL of 1.00 M HCl are required to just react with all the zinc? (You would, in reality, add much more to completely surround the penny.)

50. Assuming that the stoichiometry is 1:1, what is the concentration of an unknown acid if a 20.0-mL sample of it is neutralized with precisely 33.4 mL of 0.250 M base?

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JS Bruner

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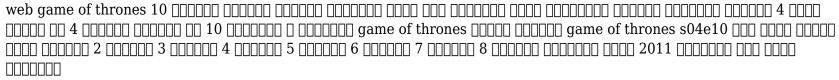
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