

Chemistry Lab Report: Standardization of Acid and Bases through Titration

2 pages | 320 words

Introduction

The laboratory method used in this experiment is titration, a fundamental technique in quantitative chemistry. Titration involves the controlled addition of a known solution (the titrant) to react with a solution of analyte until a chemical reaction is complete. This allows us to determine the concentration of the analyte accurately.

Experimental Procedure

In this experiment, we aimed to standardize a solution of sulphuric acid (H_2SO_4) using phenol red as the indicator. The following steps were performed:

1. Weighed 25 cm³ of the unknown solution
2. Titrated the solution with the standardized sulphuric acid until a color change occurred due to the indicator
3. Recorded the initial and final burette readings

Experimental Data

We conducted the titration using the specified equipment. The results of the titration are shown in the table below:

Initial Burette Reading (mL)	Final Burette Reading (mL)	Volume of Acid Used (mL)
10.2	28.6	18.4

Discussion

In the discussion, we analyzed the data we measured during the titration. By using the formula for calculating molarity (M), which is $M_1V_1 = M_2V_2$, where M_1 is the molarity of the standardized sulphuric acid, V_1 is the volume of acid used in the titration, M_2 is the molarity of the unknown solution, and V_2 is the volume of the unknown solution taken, we calculated the molarity of the unknown solution: $M_1V_1 = M_2V_2$ ($0.1 \text{ M} \times 18.4 \text{ mL} = M_2 \times 25 \text{ mL}$) $M_2 = (0.1 \text{ M} \times 18.4 \text{ mL}) / 25 \text{ mL}$ $M_2 = 0.0736 \text{ M}$ Therefore, the molarity of the unknown solution of sulphuric acid is approximately 0.0736 M. Additionally, we assessed the precision and accuracy of our results by comparing our calculated molarity to the expected value. Any sources of error were discussed, such as instrumental errors, titration endpoint determination, and experimental conditions.

Conclusion

In conclusion, the standardization of sulphuric acid was successfully achieved through titration with phenol red as the indicator. The calculated molarity of the unknown solution is approximately 0.0736 M. This information is crucial for accurate chemical analysis and future experiments in quantitative chemistry.

Standardization Of Agno3 Lab Report

Wenbin Ji



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Laboratory Manual for Chemistry: a Quantitative Approach Robert Nelson Smith, 1969 Chemical Abstracts, 1910 **Analytical Chemistry** Gary D. Christian, Purnendu K. Dasgupta, Kevin A. Schug, 2013-10-07 With the 7th Edition of Analytical Chemistry renowned chemists Purnendu Sandy Dasgupta and Kevin Schug both of the University of Texas Arlington join the author team The new edition focuses on more in depth coverage of the principles and techniques of quantitative analysis and instrumental analysis aka Analytical Chemistry The goal of the text is to provide a foundation of the analytical process tools and computational methods and resources and to illustrate with problems that bring realism to the practice and importance of analytical chemistry It is designed for undergraduate college students majoring in chemistry and in fields related to chemistry **Analytical Chemistry** Mr. Rohit Manglik, 2024-07-03 EduGorilla Publication is a trusted name in the education sector committed to empowering learners with high quality study materials and resources Specializing in competitive exams and academic support EduGorilla provides comprehensive and well structured content tailored to meet the needs of students across various streams and levels Biological Abstracts, 1928 *Agrindex*, 1991 *Glastechnische Berichte*, 1980 Vols include Patentbericht **Bulletin analytique**, 1960 **Bulletin signalétique**, 1960 **Bulletin signalétique** Centre national de la recherche scientifique (France). Centre de documentation, 1969

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