

7-1 Review and Reinforcement

Ionic Bonding

Complete each of the following sentences by filling in the appropriate word or phrase from the list below.

negative
positive
electron

noble gas
atoms
ion

ion
ion
cation

electron configuration
polyatomic
empirical formula

1. An ionic bond is an attraction between oppositely charged ____ **ions** ____.
2. Anionic bond is ____ **negative** ____ change.
3. An ionic bond is formed by losing or gaining ____ **electrons** ____.
4. The ____ **metal** ____ cation cation that cation tend to gain, lose, or share electrons in order to acquire a full set of valence electrons.
5. When sodium and chlorine form a bond bond, both have negative the electron configuration of ____ **noble gas** ____.
6. The atoms that make up ____ **polyatomic** ____ ions are bonded together by covalent bonds.
7. The ____ **crystalline** ____ method can be used to write the formula for an ionic compound.
8. The ____ **empirical formula** ____ of a compound denotes the ratio of ions in the compound. If the compound is ionic, write "ion" if it is false. Change the subscript word or words to make it true. Write your answer on the line.

____ **ionic** ____

____ **ion** ____

____ **ion** ____

____ **polyatomic** ____

____ **crystal** ____

____ **ion (polyatomic)** ____

9. Calcium bromide is composed of calcium ion (gaining two electrons).

10. A cation has a positive charge.

11. A binary ionic compound contains only one kind of cation and one kind of anion.

12. Substances have number of more than one atom.

13. The Lewis dot diagram for chlorine, a group 7A element, has six electrons.

14. A great deal of energy is produced when an ionic compound is formed.

Stars Section 1 Reinforcement Answers

James Augustus Henry Murray



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