

GAS LAWS CHEAT SHEET

Gas Variables and Common Units

Pressure (P)	Equivalent
• atm (atmosphere)	$1 \text{ atm} = 760 \text{ torr}$ $1 \text{ atm} = 101325 \text{ Pa}$ $1 \text{ bar} = 10^5 \text{ Pa}$
• torr or mmHg	
• Pa	
• bar	
Volume (V)	Equivalent
• L (dm ³)	$1 \text{ L} = 1000 \text{ mL}$ $1 \text{ kL} = 1000 \text{ L}$
• kL (m ³)	
• mL (cm ³)	
Temperature (T)	Conversion
• °C (Celsius)	$^{\circ}\text{C} = \left(\frac{9}{5} ^{\circ}\text{F} + 32 \right)$ $\text{K} = ^{\circ}\text{C} + 273.15$
• °F (Fahrenheit)	
• K (Kelvin)*	

*all calculations involving temperature must use K

Gas Laws

Ideal Gas Eq'n (IGE)	IGE, Rewritten
$PV = nRT$	$\frac{PV}{nT} = R$
Boyle's Law	Charles' Law
$P_1 V_1 = P_2 V_2$	$\frac{V_1}{T_1} = \frac{V_2}{T_2}$
Gay-Lussac's Law	Avogadro's Law
$\frac{P_1}{T_1} = \frac{P_2}{T_2}$	$\frac{V_1}{n_1} = \frac{V_2}{n_2}$
Combined Gas Law	
$\frac{P_1 V_1}{T_1} = \frac{P_2 V_2}{T_2}$	
Dalton's Law	Amagat's Law
$P_T = P_1 + P_2 + \dots + P_n$	$V_T = V_1 + V_2 + \dots + V_n$

Important Quantities

- Standard Temperature and Pressure (STP):
 0°C , 1 atm (*pre-1984*) 0°C , 1 bar (*1984 onward*)
- Volume of 1 mole of gas at STP (molar volume):
 22.414 (0°C , 1 atm) 22.711L (0°C , 1 bar)
- Ideal Gas Constant (R):
 $0.08206 \frac{\text{L} \cdot \text{atm}}{\text{mol} \cdot \text{K}}$ or $62.36 \frac{\text{L} \cdot \text{mmHg}}{\text{mol} \cdot \text{K}}$ or $8.314 \frac{\text{L} \cdot \text{kPa}}{\text{mol} \cdot \text{K}}$

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Ian Pickup



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