

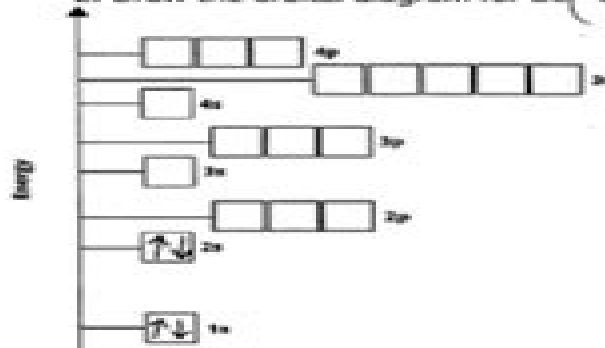
Unit 4 Review

1. Draw the orbital diagram for Ge ($32e^-$)



Name Key hr

2. Draw the orbital diagram for Be ($4e^-$)



Put the answer on the line in front of the #:

- 5 3. How many orbitals are in the d sublevel?
- 14 4. The f sublevel can hold how many electrons?
- 7 5. The f sublevel can hold how many electron pairs?
- frequency 6. The number of times a wavelength passes a given point is called its ____.
- Full 7. Is calcium's p sublevel in the 3rd energy level full or empty?
- 3 pairs 8. How many pairs of electrons does Br have in its 4th energy level?
- 2 9. How many empty orbitals does carbon have in its 2nd energy level and p sublevel?
- 3rd 10. What is the highest energy level that sodium has electrons?
- λ 11. What is the symbol for wavelength?
- ROY G BIV 12. What is the color order of the visible spectrum?
- violet 13. Which color of visible light has the highest frequency?
- Radio waves 14. What is the largest wavelength in the full electromagnetic spectrum?
- line 15. When a sample of an element is energized, like in the flame test, it releases energy in the form of a ____ spectrum.
- 2 16. The 2nd energy level has how many sublevels?
- opposite 17. Two electrons can occupy the same orbital if they have (opposite or parallel) spins.
- ground 18. When an electron is in the lowest energy orbital and energy level, the atom is in the ____ state.
- T 19. T/F. Electrons are added one at a time to the lowest energy orbitals in an atom first.
- more 20. ~~See~~ Ultraviolet radiation has more or less energy than visible light.

Unit 4 Electrons Answer Key

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A red circular graphic with a gradient, appearing as a stylized arrow or a partial circle, located to the right of the 'Mocktime Publication' text.

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