Balancing Equations Worksheet

```
10
             _{\text{H}_{2}PO_{4}} + _{\text{KOH}} \rightarrow _{\text{K}_{2}PO_{4}} + _{\text{H}_{2}O}
             K + B_0O_1 \rightarrow K_0O + B
23
30
             HO + NaOH →
                                                               NaCl + H<sub>2</sub>O
4)
             ___ Na + ___ NaNO<sub>3</sub> → ___ Na<sub>2</sub>O + ___ N<sub>2</sub>
50
             \_ C + \_ S<sub>8</sub> \rightarrow \_ CS<sub>2</sub>
60
             Na + O_2 \rightarrow Na_2O
73
             N_2 + O_2 \rightarrow N_2O_3
                    H_3PO_4 + \underline{\hspace{1cm}} Mg(OH)_2 \rightarrow \underline{\hspace{1cm}} Mg_3(PO_4)_2 + \underline{\hspace{1cm}} H_2O
80.
90
                  NaOH + ___ H<sub>2</sub>CO<sub>3</sub> → ___ Na<sub>2</sub>CO<sub>3</sub> + ___ H<sub>2</sub>O
1.00
                 KOH + HBr → KBr + H<sub>2</sub>O
             ___ Na + ___ O₂ → ___ Na₂O
113
120
             A_1(OH)_3 + H_2OO_3 \rightarrow A_2(OO_3)_3 + H_2O
             AI + S_0 \rightarrow AI_2S_0
1.30
             ___ Cs + ___ N₂ → ___ Cs₂N
1.40
150
             \_ Mg + \_ Cl<sub>2</sub> \rightarrow \_ MgCl<sub>2</sub>
160
             Rb + RbNO_3 \rightarrow Rb_2O + N_2
             C_0H_0 + O_2 \rightarrow CO_2 + H_2O
17)
             N_2 + N_2 \rightarrow NH_3
1.80
             C_{10}H_{22} + C_{2} \rightarrow C_{2} + H_{2}O
1.90
2:00
               Al(OH)_3 + HBr \rightarrow AlBr_3 + H_2O
             \underline{\hspace{0.5cm}} \underline{\hspace{0.5cm}}
211)
223
             C_1H_0 + C_2 \rightarrow CO_2 + H_2O
230
             Li + AC<sub>x</sub> → LiCl + Al
24)
             C_2H_1 + C_2 \rightarrow CO_2 + H_2O
             \_ NH<sub>4</sub>OH + \_ H<sub>3</sub>PO<sub>4</sub> \rightarrow \_ (NH<sub>4</sub>)<sub>3</sub>PO<sub>4</sub> + \_ H<sub>5</sub>O
250
             ___ Rb + ___ P → ___ Rb<sub>0</sub>P
26)
             _{-} CH<sub>4</sub> + _{-} O<sub>2</sub> \rightarrow _{-} CO<sub>2</sub> + _{-} H<sub>2</sub>O
27)
             A_1(OH)_3 + B_2O_4 \rightarrow A_2(SO_4)_3 + B_2O_4
28)
             ___ Na + ___ O₂ → ___ NaCl
290
300
             Rb + S_8 \rightarrow Rb_2S
             _{--} H<sub>3</sub>PO<sub>4</sub> + _{--} Ca(OH)<sub>2</sub> \rightarrow _{--} Ca<sub>3</sub>(PO<sub>4</sub>)<sub>2</sub> + _{--} H<sub>2</sub>O
310
             \_ NH<sub>3</sub> + \_ HQ \rightarrow \_ NH<sub>4</sub>Q
320
330
             __ Li + __ H₂O → __ LiOH + __ H₂
             _{--} Ca<sub>2</sub>(PO<sub>4</sub>)<sub>2</sub> + _{--} SiO<sub>2</sub> + _{--} C \rightarrow _{--} CaSiO<sub>3</sub> + _{--} CO + _{--}
34)
```

 $NH_3 + Q_2 \rightarrow N_2 + Q_1$

__ C + __ SO₂ → __ CS₂ + __ CO

___ FeS₂ + ___ O₂ → ___ Fe₂O₃ + ___ SO₂

35) 36)

37)

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Arthur E. Morris, H. Alan Fine, Gordon Harold Geiger

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