

11. What is the boiling point temperature of this substance (provide a temperature, not a letter)? _____

12. More/Less energy is required to boil this substance than to melt it.

Thermochemistry Terms

Fill in the blank with the correct terms. Use each term from the wordbank once.

Word Bank (yes, temperature will be used twice)

Temperature

Specific Heat
capacity

Low / Cold

Open

Heat

Temperature

Law of
Conservation of
Energy

High / Hot

Closed

1. The thing we measure when we want to determine the average kinetic energy of random motion in the particles of a substance is _____.
2. A(n) _____ system can exchange matter and energy with the surroundings.
3. The _____ is the energy needed to raise the temperature of 1 gram of a substance by one degree Celsius.
4. Energy in a phase change or chemical reaction can transform and be exchanged between the system and surroundings. However, the total energy of the universe remains the same. This is a description of _____.
5. _____ is the transfer of thermal energy that always flows from _____ to _____.
6. A(n) _____ system can exchange only energy with the surroundings.
7. When _____ increases, particles speed up or move faster.

Specific Heat/Calorimetry

13. **Application:** which contains more heat, a cup of boiling water or a large bucket of water at 55°C? Explain your answer.

$$Q = mc\Delta t$$

- Q = heat transferred (calories or joules)
- m = mass (g)
- c = specific heat (cal/g·°C) or (J/g·°C)
- Δt = change in temp. (°C)

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