Unit 1 Chemistry Study Guide

History of the Atom:

Democritus (450 BCE):

- Matter is not infinitely divisible.
- There are single particles that cannot be divided further, called atomos.

Aristotle (380 BCE):

- Matter is infinitely divisible.
- Matter is called hyle.

Antoine Lavoisier (1774):

Law of Conservation of Mass: Mass cannot be created or destroyed.

Joseph Proust (1779):

 Law of Definite Proportions: All samples of a pure compound contain the same elements in the same proportion by mass (no matter quantity).

John Dalton (1803 - 1808):

- Observed work of Lavoisier and Proust
- Led Dalton to develop the first atomic theory.
- Daltons Atomic Theory:
 - All matter is composed of extremely small particles called atoms.
 - All atoms of a given element are identical, having the same size, mass, and chemical properties. The atoms of one element are different from the atoms of all other elements.
 - Atoms cannot be subdivided, created, or destroyed.
 - Atoms to different elements can combine in simple whole number ratios to form chemical compounds.
 - o In chemical reactions, atoms are combined, separated or rearranged.
- Law of Multiple Proportions: Elements can combine in different proportions. However, the different compounds will have different properties but the ratio of atoms is always a whole number.

Billiard Ball Model: ->



allow

- J.J. Thompson (1906):
 - Cathode Ray Experiment -
 - Part 1: Thompson ran electricity through a cathode ray tube and when he put the negatively charged end of a magnet near the tube, he saw the ray of electricity

Unit 1 Chemistry Study Guide

Thomas Griffiths

Unit 1 Chemistry Study Guide:

Ccea as Chemistry Student Unit Guide Alyn McFarland, 2012 Written by a senior examiner Alyn G McFarland this CCEA AS Chemistry Student Unit Guide is the essential study companion for Unit 1 Basic Concepts in Physical and Inorganic Chemistry This full colour book includes all you need to know to prepare for your unit exam a aaa clear guidance on the content of the unit with topic summaries knowledge check questions and a quick reference indexa aaa examiner s advice throughout so you will know what to expect in the exam and will be able to demonstrate the skills required as as exam style questions with graded student responses so you can see clearly what is required toaget a better grade Basic Concepts of Chemistry, 9e Study Guide and Solutions Manual Leo J. Malone, Theodore O. Dolter, 2012-01-03 The 9th edition of Malone s Basic Concepts of Chemistry provides many new and advanced features that continue to address general chemistry topics with an emphasis on outcomes assessment New and advanced features include an objectives grid at the end of each chapter which ties the objectives to examples within the sections assessment exercises at the end each section and relevant chapter problems at the end of each chapter A new Math Check allows quick access to the needed basic skill The first chapter now includes brief introductions to several fundamental chemical concepts and Chapter Synthesis Problems have been added to the end of each chapter to bring key concepts into one encompassing problem Every concept in the text is clearly illustrated with one or more step by step examples Making it Real essays have been updated to present timely and engaging real world applications emphasizing the relevance of the material they are learning This edition continues the end of chapter Student Workshop activities to cater to the many different learning styles and to engage users in the practical aspect of the material discussed in the chapter Study Guide for Introductory Chemistry Iris Stovall,1990 **Resources in Education** ,1997

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Buttel, Sonya Ankerholz-Potts, 2004

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