Applied Science Unit 1 Principles and Applications of Science 1

2 - Answer-How many electrons can an a subshell hold?

6 - Answer-How many electrons can a p subshell hold?

10 - Answer-How many electrons can a d subshell hold?

s - Answer-Which subshells are available in the first energy level?

s and p - Answer-Which subshells are available in the second energy level?

s, p and d - Answer-Which subshells are available in the third energy level?

Orbitals must all be singly filled before they can be doubly occupied - Answer-What is Hund's rule?

Copper and chromium - Answer-Which elements do not fill the 4s subshell before the 3d subshell?

The electrostatic attraction between oppositely charged ions - Answer-Define the term ionic bond

+1 - Answer-What is the charge of an ion from group 1?

+2 - Answer-What is the charge of an ion from group 2?

-2 - Answer-What is the charge of an ion from group 6?

-1 - Answer-What is the charge of an ion from group 7?

Na loses its 2s1 electron gaining a +ve charge.

CI gains an electron in the 3p subshell gaining a -ve charge.

The opposite charges attract to form NaCl - Answer-Explain how atoms of sodium react with atoms of chlorine

Each +ve ion is surrounded by 6 -ve ions and vice versa.

Strong electrostatic attraction in every direction.

Requires a large amount of energy to break - Answer-Why do ionic bonds have such high melting points?

Size of ion and charge on ion - Answer-State two factors that affect the strength of an ionic bond

When molten or in aqueous solution - Answer-When can ionic substances conduct electricity?

Conduct electricity when molten or aqueous solution High melting/boiling points

Usually soluble in water - Answer-Describe the properties of ionic compounds

A shared pair of electrons - Answer-Define the term covalent bond

A shared pair of electrons where one atom donates both shared electrons - Answer-Define the term dative covalent bond

Unit 1 Principles Of Science Written Paper

University of Illinois (Urbana-Champaign campus)

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