

Balancing Equations Worksheet

- 1) $\text{H}_3\text{PO}_4 + \text{KOH} \rightarrow \text{K}_3\text{PO}_4 + \text{H}_2\text{O}$
- 2) $\text{K} + \text{B}_2\text{O}_3 \rightarrow \text{K}_2\text{O} + \text{B}$
- 3) $\text{HCl} + \text{NaOH} \rightarrow \text{NaCl} + \text{H}_2\text{O}$
- 4) $\text{Na} + \text{NaNO}_3 \rightarrow \text{Na}_2\text{O} + \text{N}_2$
- 5) $\text{C} + \text{S}_8 \rightarrow \text{CS}_2$
- 6) $\text{Na} + \text{O}_2 \rightarrow \text{Na}_2\text{O}$
- 7) $\text{N}_2 + \text{O}_2 \rightarrow \text{N}_2\text{O}_5$
- 8) $\text{H}_3\text{PO}_4 + \text{Mg(OH)}_2 \rightarrow \text{Mg}_3(\text{PO}_4)_2 + \text{H}_2\text{O}$
- 9) $\text{NaOH} + \text{H}_2\text{CO}_3 \rightarrow \text{Na}_2\text{CO}_3 + \text{H}_2\text{O}$
- 10) $\text{KOH} + \text{HBr} \rightarrow \text{KBr} + \text{H}_2\text{O}$
- 11) $\text{Na} + \text{O}_2 \rightarrow \text{Na}_2\text{O}$
- 12) $\text{Al(OH)}_3 + \text{H}_2\text{CO}_3 \rightarrow \text{Al}_2(\text{CO}_3)_3 + \text{H}_2\text{O}$
- 13) $\text{Al} + \text{S}_8 \rightarrow \text{Al}_2\text{S}_3$
- 14) $\text{Cs} + \text{N}_2 \rightarrow \text{Cs}_3\text{N}$
- 15) $\text{Mg} + \text{Cl}_2 \rightarrow \text{MgCl}_2$
- 16) $\text{Rb} + \text{RbNO}_3 \rightarrow \text{Rb}_2\text{O} + \text{N}_2$
- 17) $\text{C}_6\text{H}_6 + \text{O}_2 \rightarrow \text{CO}_2 + \text{H}_2\text{O}$
- 18) $\text{N}_2 + \text{H}_2 \rightarrow \text{NH}_3$
- 19) $\text{C}_{10}\text{H}_{22} + \text{O}_2 \rightarrow \text{CO}_2 + \text{H}_2\text{O}$
- 20) $\text{Al(OH)}_3 + \text{HBr} \rightarrow \text{AlBr}_3 + \text{H}_2\text{O}$
- 21) $\text{CH}_3\text{CH}_2\text{CH}_2\text{CH}_3 + \text{O}_2 \rightarrow \text{CO}_2 + \text{H}_2\text{O}$
- 22) $\text{C}_7\text{H}_8 + \text{O}_2 \rightarrow \text{CO}_2 + \text{H}_2\text{O}$
- 23) $\text{Li} + \text{AlCl}_3 \rightarrow \text{LiCl} + \text{Al}$
- 24) $\text{C}_2\text{H}_6 + \text{O}_2 \rightarrow \text{CO}_2 + \text{H}_2\text{O}$
- 25) $\text{NH}_4\text{OH} + \text{H}_3\text{PO}_4 \rightarrow (\text{NH}_4)_3\text{PO}_4 + \text{H}_2\text{O}$
- 26) $\text{Rb} + \text{P} \rightarrow \text{Rb}_3\text{P}$
- 27) $\text{CH}_4 + \text{O}_2 \rightarrow \text{CO}_2 + \text{H}_2\text{O}$
- 28) $\text{Al(OH)}_3 + \text{H}_2\text{SO}_4 \rightarrow \text{Al}_2(\text{SO}_4)_3 + \text{H}_2\text{O}$
- 29) $\text{Na} + \text{Cl}_2 \rightarrow \text{NaCl}$
- 30) $\text{Rb} + \text{S}_8 \rightarrow \text{Rb}_2\text{S}$
- 31) $\text{H}_3\text{PO}_4 + \text{Ca(OH)}_2 \rightarrow \text{Ca}_3(\text{PO}_4)_2 + \text{H}_2\text{O}$
- 32) $\text{NH}_3 + \text{HCl} \rightarrow \text{NH}_4\text{Cl}$
- 33) $\text{Li} + \text{H}_2\text{O} \rightarrow \text{LiOH} + \text{H}_2$
- 34) $\text{Ca}_3(\text{PO}_4)_2 + \text{SiO}_2 + \text{C} \rightarrow \text{CaSiO}_3 + \text{CO} + \text{P}$
- 35) $\text{NH}_3 + \text{O}_2 \rightarrow \text{N}_2 + \text{H}_2\text{O}$
- 36) $\text{FeS}_2 + \text{O}_2 \rightarrow \text{Fe}_2\text{O}_3 + \text{SO}_2$
- 37) $\text{C} + \text{SO}_2 \rightarrow \text{CS}_2 + \text{CO}$

Unit 7 Balancing Chemical Reactions Key

Darrell D. Ebbing, Albert E. Richardson

A red circular graphic with a gradient, appearing as a partial circle or a stylized arrow pointing to the right, located to the right of the authors' names.

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