

Stability

Noble gases seldom form compounds due to their stable electron configuration.

The stability of an atom, ion, or compound is related to its energy (lower energy states are more stable).

In ionic bonds- metals and nonmetals gains stability by transferring electrons.

Valence electron sharing is another way atoms can become more stable- attain a noble gas configuration (have less energy).



Covalent Bonds

Covalent Bond-

- a chemical bond that results from sharing valence electrons.
- Shared electrons are considered to be part of the outer energy levels of both atoms involved.
- majority form between atoms of nonmetallic elements.

Molecule- formed when two or more atoms bond covalently.

Diatomic Molecules-

- form when two atoms of each element share electrons (H_2 , O_2)
- The two atom molecules are more stable than the individual atoms.



Several forces act when two atoms come close together.

1. Two opposing forces act on the atoms:
 - like charged electrons
 - like charged protons
2. One atom's positive attract the other atoms electrons.

Most stable arrangement of atoms is a covalent bond- occurs at an optimal distance between nuclei. The net attraction is greater than the net repulsion.

Unbonded Pairs- lone pairs

Bonding Pair - shared electron pair.

Each covalently bonded atom equally attracts the pair of shared electrons. The shared electrons belong to each atom.

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