

QUESTION 2

(a) (i) The first ionization energy of an element is the amount of energy required to remove one electron from each atom in a mole of gaseous atoms, producing one mole of gaseous ions with a positive charge.

(ii) The electronic configuration of "Be" is $1s^2 2s^2$ and that of "B" is $1s^2 2s^2 2p^1$. All the subshells of Beryllium is filled and there is some special or extra energetic stability associated with that. This means that Beryllium is more stable than boron and hence has a higher ionization energy.

(b)

$$\frac{e \times F}{It} = \frac{M}{m}$$

$$\frac{e \times 96500}{0.12 \times 500} = \frac{48}{0.015}$$

Cross-multiplying

$$e = \frac{0.12 \times 500 \times 48}{96500 \times 0.015}$$

$$e = 1.99$$

$$e \approx 2$$

The charge is +2

c) (i) Pure alumina in molten cryolite, Na_2AlF_6

- (ii) • Availability of raw materials
- Availability of electric power

(d) (i) Paramagnetism is a property exhibited by ferrous atoms or ions having one or more unpaired electrons which are attracted into a magnetic field. e.g. Iron atom has four unpaired 3d electrons.

(d) (i) $2s^2$
I) $2s^2$

II) $2s^2$

III) $2s^2$
L
 $2s^2$

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