

## Thermodynamics worksheet

### Standard bond enthalpy

- 1) Calculate the bond energy of Cl-Cl bond from the following data



Also the bond enthalpies of C-H, C-Cl, H-Cl are 413, 326, 431 KJ/mol respectively.

- 2) Propose the structure of  $\text{C}_2\text{H}_2\text{Cl}_2$ . Calculate the change in enthalpy of the reaction.



Given that average bond enthalpies are C-C, C-H, C-Cl, C-F and C-Br are 347, 413, 326, 485, 283 kJ/mol respectively.

- 3) For the reaction  $\text{H}_2\text{O} + \text{H}_2 \rightarrow \text{H}_2\text{O}_2$  calculate the bond energy of O-O. Given bond energies of C-H, C-Cl, C-F and C-Br are 413, 326, 485, 283 and 200 kJ/mol respectively.

- 4) Calculate the bond enthalpy of H-F given that the bond enthalpies of H<sub>2</sub> and F<sub>2</sub> are 436 kJ/mol and 158 kJ/mol and enthalpy of formation for HF is -271 kJ/mol.

- 5) Calculate the bond energy of C-H bond given that the heat of formation of methane, heat of combustion of carbon, heat of dissociation of H<sub>2</sub> are -75 kJ, -393 kJ, 436 kJ/mol respectively.



- 6) Calculate the enthalpy change for the reaction  $\text{H}_2 + \text{Br}_2 \rightarrow 2\text{HBr}$ .

Given the bond enthalpies of H-H, Br-Br, H-Br is 436, 193 and 366 kJ/mol respectively.

# Unit 1 thermochemistry Review

**JL Elias**

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## **Table of Contents Unit 1thermochemistry Review**

1. Understanding the eBook Unit 1thermochemistry Review
  - The Rise of Digital Reading Unit 1thermochemistry Review
  - Advantages of eBooks Over Traditional Books
2. Identifying Unit 1thermochemistry Review
  - Exploring Different Genres
  - Considering Fiction vs. Non-Fiction
  - Determining Your Reading Goals
3. Choosing the Right eBook Platform
  - Popular eBook Platforms
  - Features to Look for in an Unit 1thermochemistry Review
  - User-Friendly Interface

4. Exploring eBook Recommendations from Unit 1thermochemistry Review
  - Personalized Recommendations
  - Unit 1thermochemistry Review User Reviews and Ratings
  - Unit 1thermochemistry Review and Bestseller Lists
5. Accessing Unit 1thermochemistry Review Free and Paid eBooks
  - Unit 1thermochemistry Review Public Domain eBooks
  - Unit 1thermochemistry Review eBook Subscription Services
  - Unit 1thermochemistry Review Budget-Friendly Options
6. Navigating Unit 1thermochemistry Review eBook Formats
  - ePub, PDF, MOBI, and More
  - Unit 1thermochemistry Review Compatibility with Devices
  - Unit 1thermochemistry Review Enhanced eBook Features
7. Enhancing Your Reading Experience
  - Adjustable Fonts and Text Sizes of Unit 1thermochemistry Review
  - Highlighting and Note-Taking Unit 1thermochemistry Review
  - Interactive Elements Unit 1thermochemistry Review
8. Staying Engaged with Unit 1thermochemistry Review
  - Joining Online Reading Communities
  - Participating in Virtual Book Clubs
  - Following Authors and Publishers Unit 1thermochemistry Review
9. Balancing eBooks and Physical Books Unit 1thermochemistry Review
  - Benefits of a Digital Library
  - Creating a Diverse Reading Collection Unit 1thermochemistry Review
10. Overcoming Reading Challenges
  - Dealing with Digital Eye Strain
  - Minimizing Distractions
  - Managing Screen Time
11. Cultivating a Reading Routine Unit 1thermochemistry Review
  - Setting Reading Goals Unit 1thermochemistry Review
  - Carving Out Dedicated Reading Time

12. Sourcing Reliable Information of Unit 1thermochemistry Review
  - Fact-Checking eBook Content of Unit 1thermochemistry Review
  - Distinguishing Credible Sources
13. Promoting Lifelong Learning
  - Utilizing eBooks for Skill Development
  - Exploring Educational eBooks
14. Embracing eBook Trends
  - Integration of Multimedia Elements
  - Interactive and Gamified eBooks

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