

Balancing Chemical Equations Worksheet

1. 2 H_2 + 1 $\text{O}_2 \rightarrow$ 2 H_2O
2. 1 N_2 + 3 $\text{H}_2 \rightarrow$ 2 NH_3
3. 1 S_8 + 12 $\text{O}_2 \rightarrow$ 8 SO_3
4. 2 N_2 + 1 $\text{O}_2 \rightarrow$ 2 N_2O
5. 2 $\text{HgO} \rightarrow$ 2 Hg + 1 O_2
6. 6 CO_2 + 6 $\text{H}_2\text{O} \rightarrow$ 1 $\text{C}_6\text{H}_{12}\text{O}_6$ + 6 O_2
7. 1 Zn + 2 $\text{HCl} \rightarrow$ 1 ZnCl_2 + 1 H_2
8. 1 SiCl_4 + 4 $\text{H}_2\text{O} \rightarrow$ 1 H_4SiO_4 + 4 HCl
9. 2 Na + 2 $\text{H}_2\text{O} \rightarrow$ 2 NaOH + 1 H_2
10. 2 $\text{H}_3\text{PO}_4 \rightarrow$ 1 $\text{H}_4\text{P}_2\text{O}_7$ + 1 H_2O
11. 1 $\text{C}_{10}\text{H}_{16}$ + 8 $\text{Cl}_2 \rightarrow$ 10 C + 16 HCl
12. 1 CO_2 + 2 $\text{NH}_3 \rightarrow$ 1 $\text{OC}(\text{NH}_2)_2$ + 1 H_2O
13. 4 Si_2H_3 + 17 $\text{O}_2 \rightarrow$ 8 SiO_2 + 6 H_2O_3
14. 2 $\text{Al}(\text{OH})_3$ + 3 $\text{H}_2\text{SO}_4 \rightarrow$ 1 $\text{Al}_2(\text{SO}_4)_3$ + 6 H_2O
15. 4 Fe + 3 $\text{O}_2 \rightarrow$ 2 Fe_2O_3
16. 1 $\text{Fe}_2(\text{SO}_4)_3$ + 6 $\text{KOH} \rightarrow$ 3 K_2SO_4 + 2 $\text{Fe}(\text{OH})_3$
17. 2 $\text{C}_7\text{H}_6\text{O}_2$ + 15 $\text{O}_2 \rightarrow$ 14 CO_2 + 6 H_2O
18. 1 H_2SO_4 + 8 $\text{HI} \rightarrow$ 1 H_2S + 4 I_2 + 4 H_2O
19. 4 FeS_2 + 11 $\text{O}_2 \rightarrow$ 2 Fe_2O_3 + 8 SO_2
20. 2 Al + 3 $\text{FeO} \rightarrow$ 1 Al_2O_3 + 3 Fe
21. 1 Fe_2O_3 + 3 $\text{H}_2 \rightarrow$ 2 Fe + 3 H_2O
22. 1 Na_2CO_3 + 2 $\text{HCl} \rightarrow$ 2 NaCl + 1 H_2O + 1 CO_2
23. 2 K + 1 $\text{Br}_2 \rightarrow$ 2 KBr
24. 1 C_7H_{16} + 11 $\text{O}_2 \rightarrow$ 7 CO_2 + 8 H_2O
25. 1 P_4 + 5 $\text{O}_2 \rightarrow$ 2 P_2O_5

Unit 6 Balancing Chemical Reactions 1 Key

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Basic Concepts of Chemistry Leo J. Malone, Theodore Dolter, 2008-12-03 Engineers who need to have a better understanding of chemistry will benefit from this accessible book. It places a stronger emphasis on outcomes assessment which is the driving force for many of the new features. Each section focuses on the development and assessment of one or two specific objectives. Within each section a specific objective is included, an anticipatory set to orient the reader, content discussion from established authors, and guided practice problems for relevant objectives. These features are followed by a set of independent practice problems. The expanded Making it Real feature showcases topics of current interest relating to the subject at hand such as chemical forensics and more medical related topics. Numerous worked examples in the text now include Analysis and Synthesis sections which allow engineers to explore concepts in greater depth and discuss outside relevance.

Chemical Reactions and Chemical Reactors George W. Roberts, 2008-03-14 Focused on the undergraduate audience, Chemical Reaction Engineering provides students with complete coverage of the fundamentals including in-depth coverage of chemical kinetics. By introducing heterogeneous catalysis early in the book, the text gives students the knowledge they need to solve real chemistry and industrial problems. An emphasis on problem solving and numerical techniques ensures students learn and practice the skills they will need later on, whether for industry or graduate work.

Physical Chemistry for Engineering and Applied Sciences Frank R. Foulkes, 2012-09-12 Physical Chemistry for Engineering and Applied Sciences is the product of over 30 years of teaching first-year Physical Chemistry as part of the Faculty of Applied Science and Engineering at the University of Toronto. Designed to be as rigorous as compatible with a first-year student's ability to understand the text, it presents detailed step-by-step derivations of the equations that permit the student to follow the underlying logic and of equal importance to appreciate any simplifying assumptions made or mathematical tricks employed. In addition to the 600 exercises and end-of-chapter problems, the text is rich in worked non-trivial examples, many of which are designed to be inspiring and thought-provoking. Step-by-step derivation of all equations enables the student to smoothly follow the derivation by sight and can be understood relatively easily by students with moderate skills and backgrounds in mathematics. Clear and accessible Physical Chemistry for Engineering and Applied Sciences includes the answers to all of the 112 worked examples, 99 exercises following many of the worked examples, and 496 end-of-chapter problems. Topics not normally seen in introductory physical chemistry textbooks: ionic reaction rates, activities and activity coefficients, or not regularly explained in much detail: electrochemistry, chemical kinetics with an eye on industrial applications. Special appendices that provide detailed explanations of basic integration and natural logarithms for students lacking a background in integral calculus. An in-depth chapter on electrochemistry in which activities and activity coefficients are used extensively as required for accurate calculations.

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Chemistry Laboratory courses at the high school or college and university level as a supplemental resource to help students master this skill It begins with simple examples from everyday life demonstrates the importance of balancing the equation addresses the role of the mole in these computations discusses different types of liquids considers the role of significant figures and culminates with the planning of syntheses There are suggestions for further reading as well as practice problems and questions to ensure mastery Begins with examples from everyday life that enable students to understand the concepts of theoretical and percent yield before applying those concepts to the laboratory Addresses the necessity of balancing the reaction equation the centrality of the mole in these calculations and the role of significant figures in reporting the answer Explains how to approach the calculations when using neat liquids or solutions The culmination of this text is the use of the same thought processes to plan the amounts of reactants needed for syntheses of desired quantities of product All of the problems in the book include detailed solutions with accompanying text to explain the answers and ancillaries also include suggestions for further reading Material Balances for Chemical Reacting Systems R.L. Cerro,B.G. Higgins,S.

Whitaker,2022-12-05 Written for use in the first course of a typical chemical engineering program Material Balances for Chemical Reacting Systems introduces and teaches students a rigorous approach to solving the types of macroscopic balance problems they will encounter as chemical engineers This first course is generally taken after students have completed their studies of calculus and vector analysis and these subjects are employed throughout this text Since courses on ordinary differential equations and linear algebra are often taken simultaneously with the first chemical engineering course these subjects are introduced as needed Teaches readers the fundamental concepts associated with macroscopic balance analysis of multicomponent reacting systems Offers a novel and scientifically correct approach to handling chemical reactions Includes an introductory approach to chemical kinetics Features many worked out problems beginning with those that can be solved by hand and ending with those that benefit from the use of computer software This textbook is aimed at undergraduate chemical engineering students but can be used as a reference for graduate students and professional chemical engineers as well as readers from environmental engineering and bioengineering The text features a solutions manual with detailed solutions for all problems as well as PowerPoint lecture slides available to adopting professors **E3**

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