

Unit 9 Worksheet: St. Michael's Concentration

Part 9: Describing Molar Concentrations

1. Sodium chloride was dissolved in water to produce a 1.5M solution. Suppose that this same solution had been added to the flask solution. This would be incorrect. **Just this value!**
Then calculate molarities (1.5 molar NaCl) for each (1.5 M) solution, if possible, and for molarities (molarities) and (molarities) of 1.5 molarities.
2. A 10.0 g solution of potassium chloride is dissolved in enough water to make 100.0 mL of solution. Calculate the molar concentration of this potassium chloride.

$$10.0 \text{ g} = \frac{1 \text{ mol}}{74.55 \text{ g/mol}} \times 100.0 \text{ mL} = 0.134 \text{ M}$$

$$100.0 \text{ mL} = \frac{1 \text{ L}}{1000 \text{ mL}} \times 0.134 \text{ M} = 0.134 \text{ M}$$

3. Find the molarity of a solution made from 25.0 g of CaCl_2 dissolved in enough water to make 0.5 L.

$$25.0 \text{ g} = \frac{1 \text{ mol}}{110.98 \text{ g/mol}} \times 0.5 \text{ L} = 0.450 \text{ M}$$

$$0.5 \text{ L} = \frac{1 \text{ L}}{0.5 \text{ L}} \times 0.450 \text{ M} = 0.900 \text{ M}$$

4. An unknown liquid solution of "molarities" and molarities is prepared by dissolving 1.00 g of sodium chloride in enough water to make a solution of 100.0 mL. Calculate the molarity of sodium in this solution.

$$1.00 \text{ g} = \frac{1 \text{ mol}}{58.44 \text{ g/mol}} \times 100.0 \text{ mL} = 0.0171 \text{ M}$$

$$100.0 \text{ mL} = \frac{1 \text{ L}}{1000 \text{ mL}} \times 0.0171 \text{ M} = 0.0171 \text{ M}$$

5. A 10.0 g solution of sodium chloride is dissolved in enough water to make 100.0 mL of solution. Calculate the molarity of sodium in this solution.

$$10.0 \text{ g NaCl} = \frac{1 \text{ mol}}{58.44 \text{ g/mol}} \times 100.0 \text{ mL} = 0.171 \text{ M}$$

Unit 9 Worksheet 3 Molar Concentration Key

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