

CHEMISTRY SAMPLE TEST 1

$$PV = nRT \rightarrow n = \frac{PV}{RT}$$

$$n = 5,003$$

1. A sample of N_2 gas occupies 4.48 L volume under standard conditions (1 atm, $0^\circ C$). What is the mass of the sample? The atomic mass of N is 14 amu.

A: 5.6 g B: 56 g C: 2.8 g D: 28 g E: 22.4 g

2. Which properties are characteristic for the nonmetals?

- ~~1~~ high electrical conductivity
~~2~~ large ionization energy
~~3~~ high electronegativity
~~4~~ low electron affinity

~~A: 1,2~~

B: 2,3

C: 3,4

D: 2,4

~~E: 1,2,3~~

3. Concerning 1.2×10^{24} CO_2 molecules, which statements are true?

The molar mass is 44 g/mol.

1. it is 12 moles.
 2. it occupies 1.2×22.4 L volume under standard conditions.
 3. it has a mass of 88 grams.
 4. it consists of 3.6×10^{24} atoms.

$$\text{mole} = \frac{\text{particle}}{\text{Avogadro's no}}$$

A: 1,3

B: 1,4

C: 2,3

D: 2,4

E: 3,4

4. Which of the following substances contain covalent bonds only?

~~1. BaCl₂(aq)~~ *metal* $\rightarrow$ ionic

2. CCl₄(l)

3. HCl(g) *H⁺ Cl⁻*

• covalent bond: two atoms share electrons
 • ionic bond: one atom donates electron to the other
 4. NH₄Br(aq) $\uparrow$ ionic
 1 non-metal + 2 metal

A: 1,2,3,4

B: 1,2,4

C: 2,3

D: 1,3

E: 2,3,4

5. Which of the following atoms are isotopes of each other?

- ~~X: 11 protons, 11 electrons, 12 neutrons~~
~~Y: 11 protons, 10 electrons, 12 neutrons~~
 Z: 11 protons, 11 electrons, 13 neutrons
 W: 12 protons, 12 electrons, 12 neutrons

Same proton no
 different neutron no

A: X and Y

B: X and W

C: Y and Z

D: Z and W

E: X and Z

6. An aqueous solution is prepared by dissolving 6 g NaOH in 250 mL final volume. What is the molar concentration of the solution?

The molar mass of NaOH is 40 g/mol

- A: 6.4 mol/L
 B: 1.6 mol/L
 C: 16 mol/L
 D: 0.16 mol/L
 E: 64 mol/L

$$M = \frac{n}{V}$$

$$\frac{6 \text{ mol}}{250 / 1000 \text{ L}}$$

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Norman Edward Gronlund



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