

Unit #10 Test Review - Stoichiometry**I. Define**

- a. What 2 measures are always conserved in a reaction? _____
- b. What is a mole ratio? _____
- c. What is a limiting reactant? _____
- d. What is actual yield? _____
- e. What is theoretical yield? _____

II. Balancing

- a. ____ Al + ____ S $\rightarrow$ ____ Al_2S_3
- b. ____ Au + ____ $\text{Cl}_2 \rightarrow$ ____ AuCl_3
- c. ____ $\text{AgNO}_3 +$ ____ K $\rightarrow$ ____ $\text{KNO}_3 +$ ____ Ag
- d. ____ $\text{C}_3\text{H}_8 +$ ____ $\text{O}_2 \rightarrow$ ____ $\text{CO}_2 +$ ____ H_2O
- e. ____ Ba + ____ $\text{O}_2 \rightarrow$ ____ BaO
- f. ____ $\text{Mn}_2\text{O}_7 \rightarrow$ ____ Mn + ____ O_2

III. Stoichiometry

- a. Given this reaction: $3 \text{O}_2 + 4 \text{Al} \rightarrow 2 \text{Al}_2\text{O}_3$

How many moles of aluminum (Al) are needed to react with 1.60 moles of oxygen (O_2)?

- b. Given this reaction: $2 \text{Li} + 1 \text{Na}_2\text{SO}_4 \rightarrow 1 \text{Li}_2\text{SO}_4 + 2 \text{Na}$

How many moles of lithium (Li) are needed to react with 4.3 moles of sodium sulfate (Na_2SO_4)?

- c. Given this reaction: $1 \text{CuO} + 1 \text{H}_2 \rightarrow 1 \text{Cu} + 1 \text{H}_2\text{O}$

If 250.0 L of hydrogen begin the above reaction at STP, what mass of copper will be formed?

Unit Stoichiometry Test Review

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