

Part I. Energy Basics, First Law of Thermodynamics, ΔH , and Stoichiometry

1. For the following processes, give the sign of q (positive or negative)

- a. If heat is given off by the system, q is negative
- b. If heat is absorbed by the system, q is positive
- c. If work is done on the system, w is positive
- d. If work is done by the system, w is negative

2. For the following processes, give the sign of q (positive or negative), when the system and surrounding are at different temperatures.

- a. Steam condenses on a cold window negative $H_2O(g) \rightarrow H_2O(l) + E$
- b. Warm milk that is placed in a refrigerator negative
- c. Ice cream that is left on the countertop positive

The internal energy of a system is : $\Delta E = q + w$

In lecture, we discussed how in chemistry, gas molecules are the system that can do PV work.

And the unit for work is in J, and the conversion between pressure, volume and energy is:

$$1 \text{ L atm} = 101.325 \text{ J}$$

3. If 900 J of heat was added to a system and 300 J of work are done on the system, what is the total internal energy?

$$\Delta E = q + w = (+900 \text{ J}) + (300 \text{ J}) = +1200 \text{ J}$$

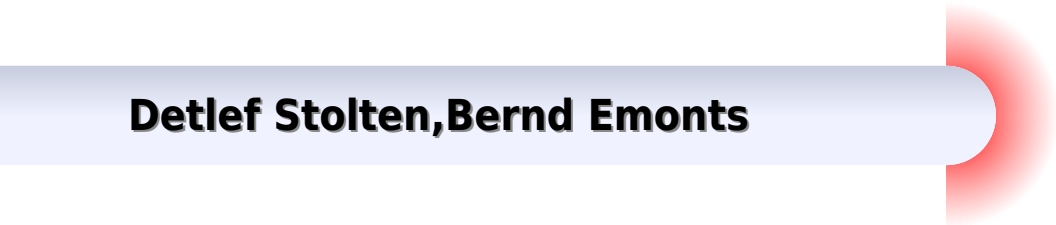
4. A system consists of one mole of gas in a cylinder (with a movable piston).

- a. If the gas absorbs 0.234 kJ of energy and 534 J of work is done on the system, what is the total internal energy (in J) of the system. Does the system expand or contract?

$$\Delta E = q + w = (+234 \text{ J}) + (+534 \text{ J}) = +768 \text{ J}$$

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