10

CHEMICAL QUANTITIES

Chapter Test A

A. Matching

Match each description in Column B to the correct term in Column A. Write the letter of the correct description on the line.

12. The mass of a mole of NaCl is its:

b. empirical formula.

a. modar mass.

Column A		Column B
	a.	the volume occupied by a mole of any gas at STP (22.4 L)
molar volume	b.	the mass in amus of an atom of an element
representative particle	5	the number of representative particles in a mole of a substance
molar mass	di.	the percent by mass of each element in a compound
atomic mass	\mathbf{e}_{-}	the mass of a mole of a substance
Wogadro's number	E.	the formula that gives the lowest whole-number ratio of the elements in a compound
mole (mol)	85-	an atom, a formula unit, or a molecule
empirical formula	b.	the mass in grams of a mole of any pure substance
percent composition	100	6.02×10^{28} representative particles of a substance
molar mass	13-	0°C and 101.3 kPa
e Choice		
	ore r	the line.
	are	in 4 molecules of isopropyl alcohol,
$4 \times 08.02 \times 10^{23}$		c. 32
	standard temperature and pressure molar volume representative particle molar mass atomic mass Avogadro's number mole (mol) empirical formula percent composition molar mass le Choice ranswer and write its letter How many hydrogen atoms	standard temperature and pressure molar volume be representative particle composition comp

c. gram atomic mass.

d. atomic mass.

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Victor M. Corman

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