

Stoichiometry "End of Chapter" IN CLASS Practice Assessment

Background Info -

- Why are the coefficients in a balanced equation necessary for stoichiometry?
Coefficients are necessary for establishing a molar ratio to convert between moles of different substances.

Skill 1 - Using a balanced equation, calculate using mole to mole ratios. Use a balanced equation, given the grams of reactant, calc. the grams of product.

- Consider the following unbalanced equation:



There are 5.56g of manganese(IV) oxide present, calculate the grams of chlorine gas.

$$5.00\text{g MnO}_2 \times \frac{1\text{mol MnO}_2}{86.94\text{g MnO}_2} \times \frac{1\text{mol Cl}_2}{1\text{mol MnO}_2} \times \frac{70.90\text{g Cl}_2}{1\text{mol Cl}_2} = 4.00\text{g Cl}_2$$

Skill 2 - After calculating theoretical yield in grams, calculate the percent yield.

- You are assigned to complete a lab. Before the reaction is complete, you calculate that a reaction should produce 4.0g Cu(s). After the lab is over you find that 2.8g Cu(s) was collected. What would the percent yield be for that reaction completed within the lab?

$$\% \text{yield} = \left[\frac{\text{actual/theoretical}}{1} \right] \times 100 \\ = \left[\frac{2.8\text{g}/4.0\text{g}}{1} \right] \times 100 = 70\%$$

Skill 3 - Given experimental data, determine the limiting reactant and the excess reactant.

- Calcium carbonate combines with gaseous hydrogen chloride to produce solid calcium chloride, gaseous carbon dioxide, and water vapor. What is the percent yield if 155g of calcium carbonate combines with 250.0g of hydrogen chloride and only 142.0g of calcium chloride is actually obtained within the lab setting? Complete and show work for all 3 steps of a stoichiometry problem.



$$155\text{g CaCO}_3 \times \frac{1\text{mol CaCO}_3}{100.09\text{g CaCO}_3} = 1.55\text{mol CaCO}_3 \quad \text{L.R.} \\ 1.55\text{mol} \\ 250\text{g HCl} \times \frac{1\text{mol HCl}}{36.45\text{g HCl}} = 6.857\text{mol HCl} \quad \text{2mol in rxn} = 3.43\text{mol}$$

$$1.55\text{mol CaCO}_3 \times \left(\frac{1\text{mol CaCl}_2}{1\text{mol CaCO}_3} \right) \times \left(\frac{110.98\text{g CaCl}_2}{1\text{mol CaCl}_2} \right) = 172.02\text{g}$$

$$\% \text{yield} = \left[\frac{\text{actual/theoretical}}{1} \right] \times 100 \\ = \left[\frac{142.0\text{g}/172.02\text{g}}{1} \right] \times 100 = 82.55\%$$

Unit 8 Stoichiometry Test Review Answers

LL Leslie

A decorative graphic consisting of a light blue horizontal bar with a rounded right end, and a red circular shape partially visible behind it.

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Table of Contents Unit 8 Stoichiometry Test Review Answers

1. Understanding the eBook Unit 8 Stoichiometry Test Review Answers
 - The Rise of Digital Reading Unit 8 Stoichiometry Test Review Answers
 - Advantages of eBooks Over Traditional Books
2. Identifying Unit 8 Stoichiometry Test Review Answers
 - Exploring Different Genres
 - Considering Fiction vs. Non-Fiction
 - Determining Your Reading Goals
3. Choosing the Right eBook Platform
 - Popular eBook Platforms
 - Features to Look for in an Unit 8 Stoichiometry Test Review Answers
 - User-Friendly Interface
4. Exploring eBook Recommendations from Unit 8 Stoichiometry Test Review Answers
 - Personalized Recommendations

- Unit 8 Stoichiometry Test Review Answers User Reviews and Ratings
- Unit 8 Stoichiometry Test Review Answers and Bestseller Lists
- 5. Accessing Unit 8 Stoichiometry Test Review Answers Free and Paid eBooks
 - Unit 8 Stoichiometry Test Review Answers Public Domain eBooks
 - Unit 8 Stoichiometry Test Review Answers eBook Subscription Services
 - Unit 8 Stoichiometry Test Review Answers Budget-Friendly Options
- 6. Navigating Unit 8 Stoichiometry Test Review Answers eBook Formats
 - ePub, PDF, MOBI, and More
 - Unit 8 Stoichiometry Test Review Answers Compatibility with Devices
 - Unit 8 Stoichiometry Test Review Answers Enhanced eBook Features
- 7. Enhancing Your Reading Experience
 - Adjustable Fonts and Text Sizes of Unit 8 Stoichiometry Test Review Answers
 - Highlighting and Note-Taking Unit 8 Stoichiometry Test Review Answers
 - Interactive Elements Unit 8 Stoichiometry Test Review Answers
- 8. Staying Engaged with Unit 8 Stoichiometry Test Review Answers
 - Joining Online Reading Communities
 - Participating in Virtual Book Clubs
 - Following Authors and Publishers Unit 8 Stoichiometry Test Review Answers
- 9. Balancing eBooks and Physical Books Unit 8 Stoichiometry Test Review Answers
 - Benefits of a Digital Library
 - Creating a Diverse Reading Collection Unit 8 Stoichiometry Test Review Answers
- 10. Overcoming Reading Challenges
 - Dealing with Digital Eye Strain
 - Minimizing Distractions
 - Managing Screen Time
- 11. Cultivating a Reading Routine Unit 8 Stoichiometry Test Review Answers
 - Setting Reading Goals Unit 8 Stoichiometry Test Review Answers
 - Carving Out Dedicated Reading Time
- 12. Sourcing Reliable Information of Unit 8 Stoichiometry Test Review Answers
 - Fact-Checking eBook Content of Unit 8 Stoichiometry Test Review Answers

- Distinguishing Credible Sources
- 13. Promoting Lifelong Learning
 - Utilizing eBooks for Skill Development
 - Exploring Educational eBooks
- 14. Embracing eBook Trends
 - Integration of Multimedia Elements
 - Interactive and Gamified eBooks

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