

Name _____

Period _____

Unit 8: Stoichiometry Test Review*Indicate whether the statement is true or false. (1-11)*

- _____ 1. *Stoichiometry* is the study of the relationship between the amount of reactants used and the amount of products made in a chemical reaction.
- _____ 2. In a balanced chemical equation, the total number of moles of all the reactants is equal to the total number of moles of all the products.
- _____ 3. In a chemical reaction represented by the general equation $A + B \rightarrow C + D$, there are four distinct mole ratios that can be written.
- _____ 4. The *mole ratio* is a comparison of how many moles of one substance are required to participate in a chemical reaction with another substance, based on the balanced chemical equation.
- _____ 5. In a chemical reaction, the reactant with the largest molar mass is the limiting reagent.
- _____ 6. The *excess reagent* refers to the additional amount of a reactant in a chemical reaction that must be added in order for the reaction to proceed.
- _____ 7. The stoichiometric relationship between any two substances in a reaction depends on the mole ratio between those substances.
- _____ 8. The *limiting reagent* limits the amount of product formed in a chemical reaction.
- _____ 9. In a balanced chemical equation, the total number mass of the reactants is equal to the total mass of the products.
- _____ 10. The percent yield is the maximum amount of product that can be produced from a given amount of reactant.
- _____ 11. If 80.0 g of sodium reacts with 72.0 g of iron (III) oxide, sodium acts as the limiting reactant.

- _____ 12. Based on the mole ratios of the substances in a chemical reaction shown, determine the correct equation for the chemical reaction?

Substances	Mole ratio
A:B	3:2
A:C	3:1
B:C	2:1

- _____ 13. Calculate the mole ratio of all the substances represented in this figure?

Unit 8 Test Stoichiometry

Rod O'Connor, Charles Mickey, Alton Hassell



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