

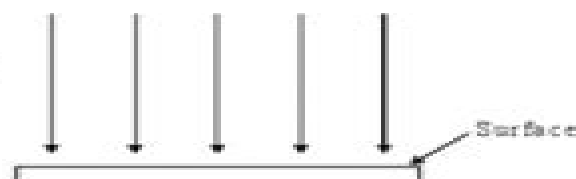
1. Some of the energy levels of the hydrogen atom are shown below. (consider all of the level jumps)



Electrons are excited to the 0.85 eV level. How many different photon frequencies will be observed in the emission spectrum of hydrogen?

- a. 3 b. 4 c. 5 **d. 6**
2. Light of wavelength 6.0×10^{-7} m is incident normally on a plane surface as shown below.

Incident light, wavelength 6.0×10^{-7} m



The light photons are absorbed by the surface.

Determine the energy for one photon of light. ($h = 6.63 \times 10^{-34}$ Js and $c = 3.00 \times 10^8$ m/s)

- a. 2.2×10^{-28} J **b. 3.3×10^{-19} J** c. 4.5×10^{-48} J d. 1.6×10^{-28} J

$$E = \frac{hc}{\lambda} = \frac{(6.63 \times 10^{-34})(3 \times 10^8)}{6.0 \times 10^{-7}} = 3.3 \times 10^{-19} \text{ J}$$

Unit Vii Quiz 1 Answers Physics

Arshad Iqbal



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