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## Chemistry Paper 3A WASSCE (SC), 2022

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### Question 1A

**A** is an aqueous solution of iodine. **B** is 0.100 mol dm<sup>-3</sup> sodium trioxothiosulphate (II).

- a. Put **B** into the burette. Pipette 20.0 cm<sup>3</sup> or 25.0 cm<sup>3</sup> of **A** into a conical flask. Add **E** from the burette until the reddish-brown colour fades to pale yellow, then add a few drops of starch indicator to obtain a dark blue solution. Continue adding **B** slowly from the burette until one drop of **B** causes the blue colour to disappear, leaving a colourless solution.

Repeat the titration to obtain concordant titre values.

Tabulate your results and calculate the average volume of **B** used.

The equation for the reaction is:



- b. From your results and the information provided, calculate the:

- i. Concentration in mol dm<sup>-3</sup> of iodine in **A**;
- ii. mass in grammes of iodine in 1.00 dm<sup>3</sup> of **A**.

[I=127.0]

[17 marks]

### Observation

Majority of the candidates responded to this question and their performance was above average.

In part (a), majority of the candidates carried out the titration experiment correctly;

In part (b), majority of the candidates calculated the concentration in mol dm<sup>-3</sup> of iodine in **A** and mass of iodine in 1.00 dm<sup>3</sup> of **A**.

The expected answers include:

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**D Siedentop**



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Chemistry Phillip John Buddle,J. S. Mackay,1970      **Chemistry Questions and Answers** Kingsley Augustine,2017-08-02 This is a chemistry book which is suitable for students in high schools or secondary schools It will also serve as a useful tool for students who are preparing for entrance examinations into colleges and universities First year students in the higher institutions taking courses in chemistry will also find this book very useful Numerous topics in chemistry have been covered by the questions contained in this handy book At the end of the question section comes the answer section which supplies all answers to the questions Therefore students are thus presented with an effective means of self assessment whereby they can determine their individual strengths and revision needs A constructive review of this chemistry handbook on objective questions will be highly appreciated from buyers so as to give ideas to others who intend to purchase a copy of this book and also to be a form of advice for the author when revising the book      Chemistry Exam

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