

Chemistry 30

Unit 4: Solutions

Assignment 3 Calculations Involving Solution and Ion Concentrations

You must clearly show all work for your calculations. Be sure to highlight your final answer – circle it, underline it, whatever. Be sure to include the units for your answer. Failure to follow these directions will result in a loss of marks.

1. What mass of sodium hydroxide is needed to make 500.0 mL of 0.5 M NaOH solution.

$$\text{Molar Mass of NaOH} = 40.0 \text{ g}\cdot\text{mol}^{-1}$$

$$g = \frac{40.0g}{\text{mol}} \times \frac{0.5\text{mol}}{L} \times \frac{0.500L}{1} = 10.0g$$

2. What volume of 0.060 M KCl solution contains 2.39 g of KCl?

$$\text{Molar mass of KCl} = 74.6 \text{ g}\cdot\text{mol}^{-1}$$

$$L = \frac{L}{0.060\text{mol}} \times \frac{\text{mol}}{74.6g} \times \frac{2.39g}{1} = 0.53L \text{ or } 530 \text{ mL}$$

3. Calculate the concentration of a solution prepared by dissolving 24.2 g of $\text{Fe}_2(\text{SO}_4)_3$ in enough water to make 250.0 mL of solution.

$$\text{Molar mass of Fe}_2(\text{SO}_4)_3 = 399.9 \text{ g}\cdot\text{mol}^{-1}$$

$$M = \frac{\text{mol}}{L} = \frac{\text{mol}}{399.9g} \times \frac{24.2g}{1} \times \frac{1}{0.250L} = 0.242M$$

Study Guide For Solution Concentration Answer Key

**Patricia A. Potter, Anne G. Perry, Kyla
C. Janzen**



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General, Organic, and Biochemistry Study Guide Marcia L. Gillette, Ira Blei, Wendy Gloffke, George Odian, 2006-01-27

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