

Now _____ Day _____ One _____

CHAPTER 9 **STUDY GUIDE FOR CONTENT MASTERY**

Covalent Bonding

Section 9.1 The Covalent Bond

In your notebook, read about the nature of covalent bonds.

Use each of the terms below just once to complete the passage.

covalent bond	molecule	sigma bond	exothermic	p bond
---------------	----------	------------	------------	--------

When sharing of electron occurs, the attachment between atoms that results is called a(n) **(1) covalent bond**. When such an attachment is formed, bond dissociation energy is released, and the process is **(2) exothermic**. When two or more atoms bond by means of electron sharing, the resulting particle is called a(n) **(3) molecule**. If the electrons shared are contained between the two atoms, the attachment is called a(n) **(4) sigma bond**. If the sharing involves the overlap of parallel orbitals, the attachment is called a(n) **(5) p bond**.

In your notebook, read about single and multiple bonds and bond strengths.

Circle the letter of the choice that best completes the statement or answers the question.

- Is what form the elements such as hydrogen, nitrogen, and oxygen normally exist?
 - as single atoms
 - as molecules containing three atoms
 - as molecules containing two atoms
 - as molecules containing four atoms
- How many electrons are shared in a double covalent bond?
 - none
 - one
 - two
 - three
- Bond length is the distance between
 - two molecules of the same substance
 - the nucleus of two attached atoms
 - the electron in two attached atoms
 - the orbitals of two attached atoms
- Which of the following relationships relating to bond length is generally correct?
 - the shorter the bond, the stronger the bond
 - the shorter the bond, the weaker the bond
 - the shorter the bond, the lower the electron in it
 - the shorter the bond, the lower the bond dissociation energy

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Now _____ Day _____ One _____

CHAPTER 9 **STUDY GUIDE FOR CONTENT MASTERY**

Section 9.2 Naming Molecules

In your notebook, read about how binary compounds and acids are named from their formulas.

For each statement below, write true or false.

- Binary molecular compounds are generally composed of a metal and a nonmetal. **false**
- The second element in the formula of a binary compound is named using the suffix -ite. **false**
- The prefix zero indicates three atoms. **false**
- The prefix four indicates six atoms. **true**
- In naming the first element in a formula, the prefix zero is not used. **true**
- For binary acids, the hydrogen part of the compound is named using the prefix hydro-. **true**
- An oxyacid contains only two elements. **false**
- If the name of the anion of an oxyacid ends in -ite, the acid name contains the suffix -ous. **false**

In your notebook, read about naming molecular compounds and oxyacids.

For each item in Column A, write the letter of the matching item in Column B.

Column A

- C_2H_6
- H_2CO_3
- HNO_3
- H_2SO_4
- H_2O_2
- H_2SO_3
- HNO_2
- H_2SO_4

Column B

- hydrobromic acid
- dinitrogen tetroxide
- carbon monoxide
- nitrous acid
- ammonia
- acetic acid
- carbonic acid
- bromic acid
- carbon dioxide

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Cynthia Spry



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Study Guide and Laboratory Manual for Physical Examination and Health Assessment, Canadian Edition-E-Book Carolyn Jarvis, Ann L. Eckhardt, 2023-02-23 NEW Preface includes an overview of the NCSBN Clinical Judgement Measurement Model PN competencies and the REx PN NEW Clinical judgement exercises equip students for success on the Next Generation NCLEX NGN including questions with an increased focus on clinical judgement Updated content throughout corresponds to the 4th Canadian edition of the textbook and incorporates the latest research and evidence informed practice

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Study Guide to Accompany Basics for Chemistry Martha Mackin, 2012-12-02 Study Guide to Accompany Basics for Chemistry is an 18 chapter text designed to be used with Basics for Chemistry textbook Each chapter contains Overview Topical Outline Skills and Common Mistakes which are all keyed to the textbook for easy cross reference The Overview section summarizes the content of the chapter and includes a comprehensive listing of terms a summary of general concepts and a list of numerical exercises while the Topical Outline provides the subtopic heads that carry the corresponding chapter and section numbers as they appear in the textbook The Fill in Multiple Choice are two sets of questions that include every concept and numerical exercise introduced in the chapter and the Skills section provides developed exercises to apply the new concepts in the chapter to particular examples The Common Mistakes section is designed to help avoid some of the errors that students make in their effort to learn chemistry while the Practical Test section includes matching and multiple choice questions that comprehensively cover almost every concept and numerical problem in the chapter After briefly dealing with an overview of chemistry this book goes on exploring the concept of matter energy measurement problem solving atom periodic table and chemical bonding These topics are followed by discussions on writing names and formulas of compounds chemical formulas and the mole chemical reactions calculations based on equations gases and the properties of a liquid The remaining chapters examine the solutions acids bases salts oxidation reduction reactions electrochemistry chemical kinetics and equilibrium and nuclear organic and biological chemistry This study guide will be of great value to chemistry teachers and students **Excel Science Study Guide, Years 9-10** Will Marchment, 2004 The book contains coverage of five major topic areas in the NSW School Certificate test Energy Force and

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