

Matter and Chemical Change Topic 1-5

1. Explain or describe how we use composition of matter to organize the world around us. Provide examples of pure substances and mixtures.

Pure substance:

- can be an element
- elements pure substances, can not be broken down further (physical and chemical wise), ex. gold
- Have a freezing and boiling point that do not change
- Compounds are another example of a pure substance. Compounds are a chemical combination of two or more elements. Can be broken down into element by heat or electricity
- Example of a Compound: H_2O = 2 hydrogen molecules + one oxygen molecule

Mix Substance:

- Matter that are not compounds or elements can be classified as either a homogenous mixture (solution) or heterogeneous mixture (mechanical mixture). These mixtures contain two or more pure substances.
- Particles of mixtures look and act different from each other.
- Mixtures have boiling, freezing and melting points that are not set.

Homogenous Mixture (Solution)

- Homo = the same
- A homogenous mixture (solution) is a physical combination of two or more chemicals mixed evenly. The solution looks the same
- Is often clear. They cannot be filtered to separate the particles. If you shine a light through, it will not scatter the light.
- Example: Water and sugar (when the sugar is fully dissolved), PEPSI, brewed coffee, air etc.

Heterogeneous Mixture

- Hetero = different
- A heterogeneous mixture (mechanical mixture) is a physical combination of two or more chemicals that are Not evenly mixed and that looks different all throughout.
- You can see the different substances that are combined together.
- Example: Chocolate chip cookie, salsa, salad dressing, pizza etc.

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