# CHAPTER 7

# THERMOCHEMISTRY SHORT QUESTION AND ANSWERS

Q.1 Define the following terms and give three examples of each.

Ans.
(i) System:

The substance which is under experiment or under observation is called as system.

Examples:

Pb(NO3)2 in decomposition of Pb(NO3)2.

(ii) Zn and CuSO4 solution, the reaction mixture in the vessel.

(iii) CaCO3 in thermal decomposition of CaCO3

(ii) Surroundings:

Everything around the system which is not a part of system is called surroundings.

For example

During the reaction between Zn and CuSO4 solution vessel and air etc are surroundings.

(iii) State function:

A macroscopic property of a system which has some definite value for initial and final state and independent of the path followed e.g.

(i) Pressure (ii) Temperature (iii) Internal energy.

Note: Heat is not a state function.

Q.2 Describe the units of energy.

Ans.

Mostly Joule and calorie are used for the measurement of energy. Calorie:

The amount of energy required to raise the temperature of one gram of water from 14.5oC to 15.5oC is called one calorie.

Joule:

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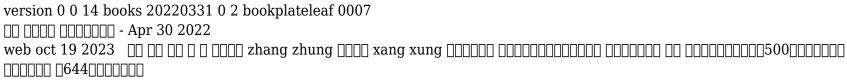
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