

Basic Five Paragraph Essay Format

Paragraph	Content	
Introduction	Sentence 1	<u>Hook</u> - One sentence designed to grab your reader's attention (quote, question, thought-provoking statement, anecdote, interesting fact, etc).
	Sentence 2-3	<u>Background Information</u> - Connect your general hook to your specific topic. What basic information should your reader know about your topic in order to understand your essay?
	Sentence 4-5	<u>Thesis</u> - One to two sentences identifying your topic, claim, and specific arguments you intend to make.
Body #1 through Body #3	Sentence 1	<u>Topic Sentence</u> : One sentence identifying what your paragraph is going to be about. This sentence should use transitions such as, first, second, third, last, also, additionally, etc.
	Sentence 2-6	<u>Details</u> : Use these sentences to prove your point using specific details from the text. The more specific you can make it, the better. Use quotes, page numbers or examples from text.
	Sentences 7-8	<u>Paragraph Summary/Transition</u> - Briefly wrap up your paragraph by putting your final thoughts on the topic at hand. Basically in your own words, what overall idea or statement do the details you supplied create? How might this topic lead into the next idea?
Conclusion	Sentence 1	<u>Wrap-Up Statement</u> : This is a one sentence overview of your topic. What is the overall point/argument/idea of your essay? You can use your original thesis to create this sentence.
	Sentence 2-3	<u>Recap Specific Ideas</u> : Use these sentences to summarize your main points from your body paragraphs. Never word your points the exact same way. Add variety by restating the ideas using new words. At no point should you introduce new ideas in your conclusion.
	Sentence 4-5	<u>Lasting Impression</u> - One to two sentences leaving your reader with a memorable idea, thought or call to action. It is clear from these sentences that your essay is done.

The Five Paragraph Essay Wizard

Jianjun Gao



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 $2 \text{ CO}_2(\text{g}) + \text{H}_2\text{O}(\text{g}) + \text{C}_2\text{H}_2(\text{g}) \rightarrow 2 \text{ H}_2\text{O}(\text{g}) + \text{C}_2\text{H}_6(\text{g})$ $\Delta H = 94.5 \text{ kJ}$
 $\text{H}_2\text{O}(\text{g}) + \text{H}_2(\text{g}) + \frac{1}{2} \text{O}_2(\text{g}) \rightarrow \text{H}_2\text{O}(\text{l})$ $\Delta H = 71.2 \text{ kJ}$
 $\text{C}_2\text{H}_6(\text{g}) + \frac{7}{2} \text{O}_2(\text{g}) \rightarrow 2 \text{ CO}_2(\text{g}) + 3 \text{ H}_2\text{O}(\text{g})$ $\Delta H = 283 \text{ kJ}$

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web chemistry 120 hess s law worksheet calculate ΔH for the reaction $\text{C}_2\text{H}_4(\text{g}) + \text{H}_2(\text{g}) \rightarrow \text{C}_2\text{H}_6(\text{g})$ from the following data
 $\text{C}_2\text{H}_4(\text{g}) + 3 \text{ O}_2(\text{g}) \rightarrow 2 \text{ CO}_2(\text{g}) + 2 \text{ H}_2\text{O}(\text{l})$ $\Delta H = 1411 \text{ kJ}$ mole
 $\text{C}_2\text{H}_6(\text{g}) + 7 \text{ O}_2(\text{g}) \rightarrow 2 \text{ CO}_2(\text{g}) + 3 \text{ H}_2\text{O}(\text{l})$ $\Delta H = 1560 \text{ kJ}$ mole
 $\text{H}_2(\text{g}) + \frac{1}{2} \text{O}_2(\text{g}) \rightarrow \text{H}_2\text{O}(\text{l})$ $\Delta H = 285 \text{ kJ}$ mole

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 hess s law example 12.1 determine ΔH for the target reaction
 $\text{C}_2\text{H}_4(\text{g}) + \text{H}_2(\text{g}) \rightarrow \text{C}_2\text{H}_6(\text{g})$
 given the following information

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 $\text{Mg}(\text{s}) + \frac{1}{2} \text{O}_2(\text{g}) \rightarrow \text{MgO}(\text{s})$ we used

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web hess s law states that the total enthalpy change for the reaction will be the sum of all those changes no matter how many different steps or stages in the reaction there are
 cohen 2016 the equations for the reactions in the experiment done are as follows
 $\text{NaOH}(\text{s}) + \text{Na}(\text{aq}) \rightarrow \text{OH}(\text{aq})$

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web thus for reaction 9 2 in which 1 mole of H_2SO_4 and 1 mole of OH^- combine to form 2 moles of H_2O $\Delta h = 55.8 \text{ kJ}$ as shown note that the enthalpy change $\Delta h_{\text{products}} - \Delta h_{\text{reactants}}$ is positive if heat is absorbed that is if $\Delta h_{\text{products}} > \Delta h_{\text{reactants}}$

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web objective the purpose of the lab is to combine two reactants in the calorimeter and measure the heats of reaction in solution eventually allowing an individual to prove or disprove hess s law introduction thermodynamics is the study of energy exchanged between the system and the surroundings

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web 3 12 hess s law in this experiment you will measure the amount of heat released in these three related exothermic reactions 1 $\text{NaOH (s)} \rightarrow \text{Na}^+ \text{(aq)} + \text{OH}^- \text{(aq)}$ ΔH_1 2 $\text{NaOH (s)} + \text{H}^+ \text{(aq)} \rightarrow \text{Na}^+ \text{(aq)} + \text{H}_2\text{O (l)}$ ΔH_2 3