# Chapter 9 (.1-.3)

# Section 9.1 (Naming Ions)

- Atoms are electrically neutral because there is the same number of protons (+) and electrons (-)
- <u>lons</u> are atoms, or groups of atoms, with a charge (either positive or negative)
  - Made by gaining or losing electrons
  - An Anion is a negative ion that gained electrons
    - Anions are usually nonmetals
    - Written like O<sup>2</sup>
  - A Cation is a positive ion that lost electrons
    - Cations are usually metals
    - Written like Ca<sup>2+</sup>
- There are two ways to name cations:
  - Using the <u>Stock System</u>, which uses roman numerals in parenthesis to indicate the numerical value
    - a. Such as Iron (II) and Iron (III)
  - Using the <u>Classical Method</u>, which uses root words yet does not give the true value (-ous, -ic)
    - a. You usually use this with cations that have more than one charge (all transition metals)
    - b. There are a few exceptions:
      - i. Silver is always Ag1+
      - ii. Candium and Zinc are always Cd2+ and Zn2+
- Since none of the anions are a transition metal, the anions always have the same charge, and you always use the <u>Classical Method</u> to name these
  - Such as Fluorine with a negative charge of one will become a fluoride ion
- <u>Polyatomic ions</u> are groups of atoms that stay together and have an overall charge, and overall name

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